

# Atomic Structure Worksheet

1. Name the three particles of the atom and their respective charges.

Particle	Charge type	Charge

2. The nucleus has a \_\_\_\_\_ charge and consist of \_\_\_\_\_ and \_\_\_\_\_.

3. Electrons are located in the \_\_\_\_\_ of an atom and carry \_\_\_\_\_ charge.

4. The number of protons in one atom of an element determines the atom's \_\_\_\_\_, and the number of electrons determines \_\_\_\_\_ of an element.

5. The atomic number tells you the number of \_\_\_\_\_ in one atom of an element. It also tells you the number of \_\_\_\_\_ in a neutral atom of that element. The atomic number gives the "identity" of an element as well as its location on the Periodic Table. No two different elements will have the \_\_\_\_\_ atomic number.

6. Give the symbol and number of protons in one atom of:

Beryllium \_\_\_\_\_ Fluorine \_\_\_\_\_ Magnesium \_\_\_\_\_ Sulfur \_\_\_\_\_  
Cobalt \_\_\_\_\_ Tin \_\_\_\_\_ Arsenic \_\_\_\_\_ Gold \_\_\_\_\_

7. Give the symbol and number of electrons in a neutral atom of:

Potassium \_\_\_\_\_ Argon \_\_\_\_\_ Calcium \_\_\_\_\_  
Chromium \_\_\_\_\_ Cadmium \_\_\_\_\_ Lead \_\_\_\_\_

8. Give the isotope symbol and number of neutrons in one atom of the following elements. Show your calculations.

Cesium - 133 \_\_\_\_\_ Iron - 54 \_\_\_\_\_  
Ruthenium - 106 \_\_\_\_\_ Nickel - 58 \_\_\_\_\_  
Polonium - 209 \_\_\_\_\_ Iodine - 135 \_\_\_\_\_  
Indium - 115 \_\_\_\_\_ Phosphorous - 31 \_\_\_\_\_

# Atomic Structure Answers Section 1

**Robert R. Redfield**

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## **Table of Contents Atomic Structure Answers Section 1**

1. Understanding the eBook Atomic Structure Answers Section 1
  - The Rise of Digital Reading Atomic Structure Answers Section 1
  - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Answers Section 1
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Atomic Structure Answers Section 1
  - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Answers Section 1
  - Personalized Recommendations
  - Atomic Structure Answers Section 1 User Reviews and Ratings
  - Atomic Structure Answers Section 1 and Bestseller Lists
5. Accessing Atomic Structure Answers Section 1 Free and Paid eBooks
  - Atomic Structure Answers Section 1 Public Domain eBooks
  - Atomic Structure Answers Section 1 eBook Subscription Services
  - Atomic Structure Answers Section 1 Budget-Friendly Options

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  - ePUB, PDF, MOBI, and More
  - Atomic Structure Answers Section 1 Compatibility with Devices
  - Atomic Structure Answers Section 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Atomic Structure Answers Section 1
  - Highlighting and Note-Taking Atomic Structure Answers Section 1
  - Interactive Elements Atomic Structure Answers Section 1
8. Staying Engaged with Atomic Structure Answers Section 1
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Atomic Structure Answers Section 1
9. Balancing eBooks and Physical Books Atomic Structure Answers Section 1
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Atomic Structure Answers Section 1
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Atomic Structure Answers Section 1
  - Setting Reading Goals Atomic Structure Answers Section 1
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Atomic Structure Answers Section 1
  - Fact-Checking eBook Content of Atomic Structure Answers Section 1
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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