

# Atomic Structure Worksheet

1. Name the three particles of the atom and their respective charges.

Particle	Charge type	Charge

2. The nucleus has a \_\_\_\_\_ charge and consist of \_\_\_\_\_ and \_\_\_\_\_.
3. Electrons are located in the \_\_\_\_\_ of an atom and carry \_\_\_\_\_ charge.
4. The number of protons in one atom of an element determines the atom's \_\_\_\_\_, and the number of electrons determines \_\_\_\_\_ of an element.
5. The atomic number tells you the number of \_\_\_\_\_ in one atom of an element. It also tells you the number of \_\_\_\_\_ in a neutral atom of that element. The atomic number gives the "identity" of an element as well as its location on the Periodic Table. No two different elements will have the \_\_\_\_\_ atomic number.
6. Give the symbol and number of protons in one atom of:
- Beryllium \_\_\_\_\_ Fluorine \_\_\_\_\_ Magnesium \_\_\_\_\_ Sulfur \_\_\_\_\_  
Cobalt \_\_\_\_\_ Tin \_\_\_\_\_ Arsenic \_\_\_\_\_ Gold \_\_\_\_\_
7. Give the symbol and number of electrons in a neutral atom of:
- Potassium \_\_\_\_\_ Argon \_\_\_\_\_ Calcium \_\_\_\_\_  
Chromium \_\_\_\_\_ Cadmium \_\_\_\_\_ Lead \_\_\_\_\_
8. Give the isotope symbol and number of neutrons in one atom of the following elements. Show your calculations.

Cesium - 133 \_\_\_\_\_

Iron - 54 \_\_\_\_\_

Ruthenium - 106 \_\_\_\_\_

Nickel - 58 \_\_\_\_\_

Polonium - 209 \_\_\_\_\_

Iodine - 135 \_\_\_\_\_

Indium - 115 \_\_\_\_\_

Phosphorous - 31 \_\_\_\_\_

# Atomic Structure Answers Section 1

**Robert R. Redfield**



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