

Science

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(Chapter - 2) (Acids, Bases and Salts) (Exercises)
(Class 10)

Question 1:

A solution turns red litmus blue, its pH is likely to be

(a) 1 (b) 4 (c) 5 (d) 10

Answer 1:

(d) 10

Basic solution has a pH value more than 7. Since the solution turns red litmus blue, its pH is likely to be 10.

Question 2:

A solution reacts with crushed egg-shells to give a gas that turns lime-water milky. The solution contains

(a) NaCl (b) HCl (c) LiCl (d) KCl

Answer 2:

(b) HCl

The crushed egg-shells contains calcium carbonate [CaCO_3]. When HCl solution reacts with this crushed egg-shells and produces carbon dioxide which turns lime - water milky.



Question 3:

10 mL of a solution of NaOH is found to be completely neutralised by 8 mL of a given solution of HCl . If we take 20 mL of the same solution of NaOH , the amount HCl solution (the same solution as before) required to neutralise it will be

(a) 4 ml (b) 8 ml (c) 12 ml (d) 16 ml

Answer 3:

(d) 16 ml

Because for the neutralisation, the ratio of the acid and base should be same.

Question 4:

Which one of the following types of medicines is used for treating indigestion?

(a) Antibiotic (b) Analgesic (c) Antacid (d) Antiseptic

Answer 4:

(c) Antacid

Because antacid reduced the acidity (produced due to excess release of HCl) in the stomach.

Question 5:

Write word equations and then balanced equations for the reaction taking place when -

(a) dilute sulphuric acid reacts with zinc granules.
(b) dilute hydrochloric acid reacts with magnesium ribbon.
(c) dilute sulphuric acid reacts with aluminium powder.
(d) dilute hydrochloric acid reacts with iron filings.

Answer 5:

(a) Sulphuric acid + Zinc \longrightarrow Zinc sulphate + Hydrogen



(b) Hydrochloric acid + Magnesium \longrightarrow Magnesium chloride + Hydrogen



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Table of Contents 10 Class Science Notes Chapterwise

1. Understanding the eBook 10 Class Science Notes Chapterwise
 - The Rise of Digital Reading 10 Class Science Notes Chapterwise
 - Advantages of eBooks Over Traditional Books
2. Identifying 10 Class Science Notes Chapterwise
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 10 Class Science Notes Chapterwise
 - User-Friendly Interface
4. Exploring eBook Recommendations from 10 Class Science Notes Chapterwise
 - Personalized Recommendations
 - 10 Class Science Notes Chapterwise User Reviews and Ratings
 - 10 Class Science Notes Chapterwise and Bestseller Lists

5. Accessing 10 Class Science Notes Chapterwise Free and Paid eBooks
 - 10 Class Science Notes Chapterwise Public Domain eBooks
 - 10 Class Science Notes Chapterwise eBook Subscription Services
 - 10 Class Science Notes Chapterwise Budget-Friendly Options
6. Navigating 10 Class Science Notes Chapterwise eBook Formats
 - ePUB, PDF, MOBI, and More
 - 10 Class Science Notes Chapterwise Compatibility with Devices
 - 10 Class Science Notes Chapterwise Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 10 Class Science Notes Chapterwise
 - Highlighting and Note-Taking 10 Class Science Notes Chapterwise
 - Interactive Elements 10 Class Science Notes Chapterwise
8. Staying Engaged with 10 Class Science Notes Chapterwise
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 10 Class Science Notes Chapterwise
9. Balancing eBooks and Physical Books 10 Class Science Notes Chapterwise
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 10 Class Science Notes Chapterwise
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 10 Class Science Notes Chapterwise
 - Setting Reading Goals 10 Class Science Notes Chapterwise
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 10 Class Science Notes Chapterwise
 - Fact-Checking eBook Content of 10 Class Science Notes Chapterwise
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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