



Real Numbers

$$R = \{-\infty, +\infty\}$$

Whole Numbers

$$W = \{0, 1, 2, 3, 4, \dots\}$$

Natural Numbers also Counting Numbers

$$N = \{1, 2, 3, 4, \dots\}$$

Integers

$$Z = \{\dots, -2, -1, 0, 1, 2, \dots\}$$

The imaginary unit i is the square root of -1 .

$$\sqrt{-1} = i$$

$$i^2 = -1$$

An imaginary number is a number that contains the imaginary unit i . $\sqrt{-a} = i\sqrt{a}$ where $a \geq 0$.

complex number $a + bi$

combination of a real number and an imaginary number

The conjugate of $a + bi$ is $a - bi$

$$(a + bi)(a - bi) = a^2 + b^2$$

$$\sqrt[n]{x} = x^{\frac{1}{n}}$$

$$x^{\frac{m}{n}} = \sqrt[n]{x^m} = (\sqrt[n]{x})^m$$

$$\sqrt[n]{a^n} = a$$

$$\sqrt[n]{ab} = \sqrt[n]{a} \cdot \sqrt[n]{b}$$

$$\sqrt[n]{\frac{a}{b}} = \frac{\sqrt[n]{a}}{\sqrt[n]{b}}$$

Compound interest is interest that applies not only to the principal, but also to previously accumulated interest

$$A = P \left(1 + \frac{r}{n}\right)^{nt}$$

amount after t years principal annual interest rate n compounds per year t years

Quadratic Formula

$$x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$$

$$\log_x(y) = z$$

base
the number we multiply by itself

logarithm
the number of multiplications

If $y = \log_b x$, then $x = b^y$

Common log

The base is 10

$$\log x = \log_{10} x$$

$$\log_b a = \frac{\log a}{\log b}$$

$$\log_b b = 1$$

$$\log_b 1 = 0$$

Natural log $e = 2.71828\dots$

$$\ln x = \log_e x$$

$$\ln 1 = 0 \quad \ln e = 1 \quad \ln e^x = x$$

$\ln x$ and e^x are inverse functions.

$$\ln e = 1$$

$$e^{\ln 3} = 3$$

$$\ln e^x = x$$

$$e^{\ln x} = x$$

$$e \ln x = x$$

$$b^x = b^y \rightarrow x = y$$

where $b > 0$ and $b \neq 1$

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time the total pressure $p_a = x_a p_{\text{total}}$ mole fraction $x_a = \frac{\text{moles of substance } a}{\text{total moles}}$
a is often denoted by x_a and

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web 1 in the given mixture the sum of all the mole fractions is equal to one $x_a + x_b + \dots = 1$ the multiplication of the mole fraction
by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction
mole fraction is not dependent on the temperature

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present is equal to one $x_a + x_b = 1$ when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is
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ways to represent the concentration of the solution it is defined as the number of molecules of component a divided by the
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mass $C_6H_6 = 78 \text{ g/mol}$ hence C_6H_6 's number of moles $\frac{30}{78} \text{ mol}$ molar mass of carbon tet $CCl_4 = 154 \text{ g/mol}$

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g of water calculate the molarity and molality of the solution given $n(H) = 1$ $n(C) = 12$ $n(O) = 16$ given mass of solute urea 0.211 g mass
of solvent water 100 g to find molarity of solution solution

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