

# Stoichiometry Worksheet #1

1. Aluminum chloride,  $\text{AlCl}_3$ , is used as a catalyst in various industrial reactions. It is prepared from hydrogen chloride gas and aluminum metal shavings.



1A. How many moles of  $\text{AlCl}_3$  can be prepared from 3.5 moles of hydrogen chloride gas with an excess of aluminum?

$$3.5 \text{ mol HCl} \times \frac{2 \text{ mol AlCl}_3}{6 \text{ mol HCl}} = 1.2 \text{ mol AlCl}_3$$

Ans: 1.2 mol  $\text{AlCl}_3$

1B. How many moles of hydrogen gas would be produced from the use of 8.5 moles of aluminum with an excess of hydrogen chloride?

Ans: 13 mol  $\text{H}_2$

$$8.5 \text{ mol Al} \times \frac{3 \text{ mol H}_2}{2 \text{ mol Al}} = 13 \text{ mol H}_2$$

2. When dinitrogen pentoxide,  $\text{N}_2\text{O}_5$ , a white solid, is heated, it decomposes to nitrogen dioxide and oxygen.



2A. How many moles of nitrogen dioxide can be formed from the decomposition of 1.25 g of  $\text{N}_2\text{O}_5$ ?

$$1.25 \text{ g N}_2\text{O}_5 \times \frac{1 \text{ mol N}_2\text{O}_5}{108.02 \text{ g N}_2\text{O}_5} \times \frac{4 \text{ mol NO}_2}{2 \text{ mol N}_2\text{O}_5} = 0.0231 \text{ mol NO}_2$$

Ans: 0.0231 mol  $\text{NO}_2$

2B. How many grams of oxygen can be formed from the decomposition of 2.3 g of  $\text{N}_2\text{O}_5$ ?

$$2.3 \text{ g N}_2\text{O}_5 \times \frac{1 \text{ mol N}_2\text{O}_5}{108.02 \text{ g N}_2\text{O}_5} \times \frac{1 \text{ mol O}_2}{2 \text{ mol N}_2\text{O}_5} \times \frac{32.00 \text{ g O}_2}{1 \text{ mol O}_2} = 0.34 \text{ g O}_2$$

Ans: 0.34 g  $\text{O}_2$

3. Chlorine is prepared from sodium chloride by electrochemical decomposition. Formerly chlorine was produced by heating hydrochloric acid with pyrolusite [manganese dioxide or manganese(IV) oxide,  $\text{MnO}_2$ ], a common manganese ore. Small amounts of chlorine may be prepared in the laboratory by the same reaction.



3A. How many grams of HCl react with 5.00 g of  $\text{MnO}_2$ , according to the equation?

$$5.00 \text{ g MnO}_2 \times \frac{1 \text{ mol MnO}_2}{86.94 \text{ g MnO}_2} \times \frac{4 \text{ mol HCl}}{1 \text{ mol MnO}_2} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 8.39 \text{ g HCl}$$

Ans: 8.39 g HCl

3B. If a chemist wanted to prepare 100. g of chlorine, how many grams of  $\text{MnO}_2$  are needed, assuming there is more than enough hydrochloric acid?

Ans: 123 g  $\text{MnO}_2$

$$100. \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.90 \text{ g Cl}_2} \times \frac{1 \text{ mol MnO}_2}{1 \text{ mol Cl}_2} \times \frac{86.94 \text{ g MnO}_2}{1 \text{ mol MnO}_2} = 123 \text{ g MnO}_2$$

3C. How many molecules of water are produced from the reaction of 5.0 g of HCl?

Ans:  $4.1 \times 10^{23}$  molec.  $\text{H}_2\text{O}$

$$5.0 \text{ g HCl} \times \frac{1 \text{ mol HCl}}{36.46 \text{ g HCl}} \times \frac{2 \text{ mol H}_2\text{O}}{4 \text{ mol HCl}} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 4.1 \times 10^{23} \text{ molecules H}_2\text{O}$$

4. Sodium is a soft, reactive metal that instantly reacts with water to give hydrogen gas and a solution of sodium hydroxide,



4A. How many grams of sodium metal are needed to give 7.81 g of hydrogen by this reaction?

Ans: 178 g Na

$$7.81 \text{ g H}_2 \times \frac{1 \text{ mol H}_2}{2.02 \text{ g H}_2} \times \frac{2 \text{ mol Na}}{1 \text{ mol H}_2} \times \frac{22.99 \text{ g Na}}{1 \text{ mol Na}} = 178 \text{ g Na}$$

4B. How many sodium atoms are needed to react with  $1.25 \times 10^{24}$  molecules of water?

Ans:  $1.25 \times 10^{24}$  atoms Na

$$1.25 \times 10^{24} \text{ molecules H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{6.02 \times 10^{23} \text{ molecules H}_2\text{O}} \times \frac{2 \text{ mol Na}}{2 \text{ mol H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ atoms Na}}{1 \text{ mol Na}} =$$

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**JS Bruner**



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web the presentation includes a clause by clause review of the iso 45001 2018 standard which replaces ohsas 18001 2007

provide helpful explanations and some examples of how the various requirements are met this training is well suited for client organizations looking to migrate from ohsas 18001 2007 or for new client organizations pursuing

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web ohsas 18001 2007 was developed in response to demand for a recognizable health and safety management system standard that could be externally assessed and certified with oh s management extending to global organizations and complex supply chains an international standard was needed to ensure a consistent approach for workers worldwide

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web mar 12 2018 while iso 45001 draws on ohsas 18001 the former benchmark for oh s it is a new and distinct standard not a revision or update and is due to be phased in gradually over the next three years organizations will therefore need to revise their current thinking and work practices in order to maintain organizational compliance

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web configure the following on r1 minimum password length is 10 characters encrypt plaintext passwords privileged exec mode secret password is ciscoenapa55 console line

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the router central isp cluster and the web server are completely configured you must create a new ipv4 addressing scheme that will accommodate 4 subnets using the 192 168 0 0 24 network the it department requires 25 hosts the sales department needs 50 hosts the subnet for the rest of the staff see more

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web jan 2 2020 disable dns lookup assign ciscoenpa55 as the encrypted privileged exec mode password assign ciscoconpa55 as the console password and enable

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web feb 11 2019 use the ip address of the vpn server located in the addressing table and ping the vpn server connect the client to site vpn with user sally and password

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web this activity allows you to practice a variety of skills including configuring vlans ppp with chap static and default routing using ipv4 and ipv6 due to the sheer number of

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