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Properties of Water Worksheet

1. The hydrogen and oxygen atoms are held together by Covalent bonds.
2. The electrons are not shared equally creating a polar molecule.
3. The polarity of water allows it to combine with most substances.
Because of this, it is referred to as the universal solvent.
4. Water molecules stick to other water molecules. This property is called cohesion.
5. Hydrogen bonds form between adjacent water molecules because the positive charged hydrogen end of one water molecule attracts the negative charged oxygen end of another water molecule.
6. Water molecules stick to other materials due to its polar nature. This property is called adhesion.
7. Surface tension creates the skin-like surface formed due to the polar nature of water.
8. Water has a high specific heat, it can absorb a large amount of heat before a change in temperature takes place.
9. Evaporation is when water changes from a liquid to a gas.
10. Condensation is when water changes from a gas to a liquid.
11. Substances that are attracted to water are referred to as hydrophilic and are considered to be water loving.
12. Substances that are not attracted to water are referred to as hydrophobic and are considered to be water fearing.

Word Bank:

polar
cohesion
negative
condensation

hydrophilic
adhesion
high specific heat
hydrophobic

universal solvent
combine

evaporation
covalent
surface tension
positive

13. Draw and label a water molecule.



Identify the positive (+) and negative ends (-).

14. Give the chemical formula for water.



Assessing Water Quality Lab Answer Key

RD Boyd

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