

Acids & Bases: Assessment Review

Name _____

1. **Identify** – Which substance will be acidic, basic, nonreactive acid, and nonreactive base in each group? – Please initial them.

2. **Write** a chemical equation showing how HClO_4 behaves as an acid when dissolved in water. Label the acid, base, conjugate acid & conjugate base.

3. **Write** a chemical equation showing how ammonia (NH_3) behaves as a base when dissolved in water. Label the acid, base, conjugate acid & conjugate base.

4. **Identify** – Which substance will be acidic, basic, nonreactive acidic, basic, or neutral? (pH from pH 0-14) (a neutral pH check calculation.)

5. **Examine** the calculation flowchart on the diagram provided. (be careful with these numbers!)



6. **What** is $[\text{OH}^-]$ in a solution that $[\text{H}^+] = 7.3 \times 10^{-7}$ M? Is the solution acidic, basic, or neutral?

7. **Calculate** $[\text{H}^+]$ in a solution of $[\text{OH}^-] = 5.2 \times 10^{-6}$ M. Is the solution acidic, basic, or neutral?

8. **Calculate** the $[\text{H}^+]$ in a solution that shows a pH of 1.1. Is the solution acidic, basic, or neutral?

9. **What** is the pOH of a solution of $[\text{OH}^-] = 1.2 \times 10^{-9}$ M? Is the solution acidic, basic, or neutral?

Assessment Acids And Bases Answer Key

RC Schank

Assessment Acids And Bases Answer Key:

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