

Assignment #3 – Compound Names and Formulas
Single and Multi-valent ions

A. Write the correct formula for the following compounds.

- | | |
|--|---|
| 1. gallium oxide <u>Ga_2O_3</u> | 11. silicon oxide <u>SiO_2</u> |
| 2. mercury (I) chloride <u>$HgCl$</u> | 12. aluminum bromide <u>$AlBr_3$</u> |
| 3. tin (IV) bromide <u>$SnBr_4$</u> | 13. tin (II) sulphide <u>SnS</u> |
| 4. indium oxide <u>In_2O_3</u> | 14. germanium bromide <u>$GeBr_4$</u> |
| 5. cadmium sulphide <u>CdS</u> | 15. nickel (II) oxide <u>NiO</u> |
| 6. magnesium oxide <u>MgO</u> | 16. strontium astatide <u>$SrAt_2$</u> |
| 7. potassium nitride <u>K_3N</u> | 17. aluminum oxide <u>Al_2O_3</u> |
| 8. iron (III) sulphide <u>Fe_2S_3</u> | 18. copper (II) oxide <u>CuO</u> |
| 9. gold (III) bromide <u>$AuBr_3$</u> | 19. lawrencium fluoride <u>LrF_3</u> |
| 10. zinc oxide <u>ZnO</u> | 20. actinium oxide <u>Ac_2O_3</u> |

B. Name the following compounds correctly. Only include the Roman Numeral in the name if the first element has 2 or more combining capacities.

- | | |
|--|---|
| 1. $CaBr_2$ <u>Calcium Bromide</u> | 9. $AsCl_5$ <u>Arsenic (V) Chloride</u> |
| 2. K_2O <u>Potassium Oxide</u> | 10. Eu_2O_3 <u>Europium (III) Oxide</u> |
| 3. Fe_2O_3 <u>Iron (III) Oxide</u> | 11. $NdBr_3$ <u>Neodymium Bromide</u> |
| 4. ZnF_2 <u>Zinc Fluoride</u> | 12. Cu_3N_2 <u>Copper (II) Nitride</u> |
| 5. $CuCl$ <u>Copper (I) Chloride</u> | 13. CoS <u>Cobalt (II) Sulphide</u> |
| 6. Ni_2O_3 <u>Nickel (III) Oxide</u> | 14. ZrO_2 <u>Zirconium Oxide</u> |
| 7. $AuCl$ <u>Gold (I) Chloride</u> | 15. Li_2O <u>Lithium Oxide</u> |
| 8. $HgBr$ <u>Mercury (I) Bromide</u> | 16. NaH <u>Sodium Hydride</u> |

Answer Key To Covalent Compound Naming

M Lipman



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Table of Contents Answer Key To Covalent Compound Naming

1. Understanding the eBook Answer Key To Covalent Compound Naming
 - The Rise of Digital Reading Answer Key To Covalent Compound Naming
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Covalent Compound Naming
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To Covalent Compound Naming
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Covalent Compound Naming
 - Personalized Recommendations
 - Answer Key To Covalent Compound Naming User Reviews and Ratings
 - Answer Key To Covalent Compound Naming and Bestseller Lists
5. Accessing Answer Key To Covalent Compound Naming Free and Paid eBooks
 - Answer Key To Covalent Compound Naming Public Domain eBooks
 - Answer Key To Covalent Compound Naming eBook Subscription Services
 - Answer Key To Covalent Compound Naming Budget-Friendly Options
6. Navigating Answer Key To Covalent Compound Naming eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key To Covalent Compound Naming Compatibility with Devices
 - Answer Key To Covalent Compound Naming Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To Covalent Compound Naming
 - Highlighting and Note-Taking Answer Key To Covalent Compound Naming
 - Interactive Elements Answer Key To Covalent Compound Naming

8. Staying Engaged with Answer Key To Covalent Compound Naming
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To Covalent Compound Naming
9. Balancing eBooks and Physical Books Answer Key To Covalent Compound Naming
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To Covalent Compound Naming
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key To Covalent Compound Naming
 - Setting Reading Goals Answer Key To Covalent Compound Naming
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key To Covalent Compound Naming
 - Fact-Checking eBook Content of Answer Key To Covalent Compound Naming
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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