

1999 Advanced Placement Chemistry Exam

Section I: Multiple Choice

Time — 1 hour and 30 minutes

NO CALCULATORS MAY BE USED WITH SECTION I.

Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified.

Throughout the test the following symbols have the definitions specified unless otherwise noted.

T — temperature
P — pressure
V — volume
S — entropy
H — enthalpy
G — free energy
R — molar gas constant
n — number of moles
M — molar

mol — molal
L, mL — liter(s), milliliter(s)
g — gram(s)
nm — nanometer(s)
atm — atmospheric(s)
J, kJ — joule(s), kilojoule(s)
V — volt(s)
mol — mole(s)

Part A

Directions: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then fill in the corresponding oval on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–4 refer to the following types of energy.

- (A) Activation energy
- (B) Free energy
- (C) Ionization energy
- (D) Kinetic energy
- (E) Lattice energy

1. The energy required to convert a ground-state atom in the gas phase to a gaseous positive ion
2. The energy change that occurs in the conversion of an ionic solid to widely separated gaseous ions
3. The energy in a chemical or physical change that is available to do useful work
4. The energy required to form the transition state in a chemical reaction

Questions 5–8 refer to atoms for which the occupied atomic orbitals are shown below.

- (A) $1s \underline{\hspace{2em}} 2s \underline{\hspace{2em}}$
- (B) $1s \underline{\hspace{2em}} 2s \underline{\hspace{2em}}$
- (C) $1s \underline{\hspace{2em}} 2s \underline{\hspace{2em}} 2p \underline{\hspace{2em}} \underline{\hspace{2em}} \underline{\hspace{2em}}$
- (D) $1s \underline{\hspace{2em}} 2s \underline{\hspace{2em}} 2p \underline{\hspace{2em}} \underline{\hspace{2em}} \underline{\hspace{2em}}$
- (E) $[\text{Ar}] 4s \underline{\hspace{2em}} 3d \underline{\hspace{2em}} \underline{\hspace{2em}} \underline{\hspace{2em}}$

5. Represents an atom that is chemically unreactive
6. Represents an atom in an excited state
7. Represents an atom that has four valence electrons
8. Represents an atom of a transition metal

Questions 9–12 refer to aqueous solutions containing 1:1 mole ratios of the following pairs of substances. Assume all concentrations are 1 M.

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Norman Reid,2021-01-18 Chemistry is often seen as a difficult subject to understand This book focusses on the triangle model that Alex H Johnstone developed in the early 1980s Originally conceived in the context of making chemistry more accessible to a wider range of learners the model has been applied in almost every area of education in chemistry at all stages of learning In looking at why chemistry is difficult there are two central questions Firstly does the problem relate to the nature of chemistry and secondly does it relate to the way humans gain understanding Both were found to be important and the answers to the two question were found to be connected The triangle model arose from sustained research into human learning The central finding from research is the critical role of working memory and the model rationalises so much evidence from chemistry education research as well as the repeated experiences of teachers of chemistry at all levels In order to understand chemistry it is essential to develop sound mental models of molecular reality It generates major implications for the way a chemistry curriculum should be constructed and the processes of teaching and learning in chemistry when the goal is focussed on understanding the key ideas Some of these implications are developed and pointers offered to more successful ways forward The power of the Johnstone Triangle lies in the way it offers clear directions for all involved in chemistry education It is hoped that this book will prove helpful to all involved in sharing the exciting story of the way humans have come to understand the molecular world one of the great examples of great human endeavour **The Qualified Student**

Harold S. Wechsler,2017-09-29 In The Qualified Student Harold S Wechsler focuses on methods of student selection used by institutions of higher education in the United States More specifically he discusses the way that college and university reformers employed those methods to introduce higher education into a broader cross section of America by extending access to an increased number of students from nontraditional backgrounds Implicit in much of this book is an underlying social and ethical question How legitimate was and is higher education's regulation of social mobility Public concern over colleges and universities practices became inevitable once they became regulators between social classes The challenging of colleges admissions policies in the courts augments similar concerns that have been present in legislatures for decades The volume is divided into three main sections Prerequisites Columbia and the Selective Function and Implications It focuses mainly on four universities The University of Michigan Columbia University the University of Chicago and the City University of New York Wechsler maintains that unlike other universities these institutions were pacesetters they did not adopt a new policy simply because some other college had already adopted it A new introduction brings the book originally published in 1977 up to date and demonstrates its continuing importance in today's academic world of selective admissions

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