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2  $CO_2(g)$   $\Delta h = 20\text{ kJ}$   
5  $C_2H_2(g)$   $\Delta h = 202\text{ kJ}$   
2  $H_2(g)$   $\Delta h = 2\text{ kJ}$   
6  $C_2H_6(g)$   $\Delta h = 94.5\text{ kJ}$   
2  $H_2O(g)$   $\Delta h = 112\text{ kJ}$   
 $\frac{1}{2}O_2(g)$   $\Delta h = 71.2\text{ kJ}$   
7  $C_2H_6(g)$   $\Delta h = 202\text{ kJ}$

co<sub>2</sub> g 3 h 2 o g  $\Delta h$  283 kJ g 1 c<sub>2</sub>h<sub>6</sub> g

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web thus for reaction 9 2 in which 1 mole of  $h_3o$  and 1 mole of  $oh$  combine to form 2 moles of  $h_2o$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h$  is positive if heat is absorbed that is if  $h_{products}$  is greater than  $h_{reactants}$

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