

## PART-III : MATHEMATICS

### SECTION – 1 (Maximum marks : 24)

- This section contains **EIGHT (08)** questions.
- The answer to each question is a **NUMERICAL VALUE**.
- For each question, enter the correct numerical value of the answer using the mouse and the on-screen virtual numeric keypad in the place designated to enter the answer. If the numerical value has more than two decimal places, truncate/round-off the value to **TWO** decimal places.
- Answer to each question will be evaluated according to the following marking scheme:

**Full Marks** : +3    **ONLY if the correct numerical value is entered;**

**Zero Marks** : 0    **In all other cases.**

1. Considering only the principal values of the inverse trigonometric functions, the value of

$$\frac{3}{2} \cos^{-1} \sqrt{\frac{2}{2+\pi^2}} + \frac{1}{4} \sin^{-1} \frac{2\sqrt{2}\pi}{2+\pi^2} + \tan^{-1} \frac{\sqrt{2}}{\pi}$$

is \_\_\_\_\_.

Answer (= 2.36)

$$\text{Sol. } \frac{3}{2} \tan^{-1} \frac{\pi}{\sqrt{2}} + \frac{1}{4} \tan^{-1} \left( \frac{2\sqrt{2}\pi}{\pi^2 - 2} \right) + \tan^{-1} \frac{\sqrt{2}}{\pi}$$

$$= \frac{\pi}{2} + \frac{1}{2} \tan^{-1} \frac{\pi}{\sqrt{2}} - \frac{1}{4} \tan^{-1} \left( \frac{2\sqrt{2}\pi}{2 - \pi^2} \right)$$

$$= \frac{\pi}{2} + \frac{1}{2} \tan^{-1} \left( \frac{\pi}{\sqrt{2}} \right) - \frac{1}{4} \tan^{-1} \left( \frac{2 \cdot \left( \frac{\pi}{\sqrt{2}} \right)}{1 - \left( \frac{\pi}{\sqrt{2}} \right)^2} \right)$$

$$= \frac{\pi}{2} + \frac{1}{2} \tan^{-1} \left( \frac{\pi}{\sqrt{2}} \right) - \frac{1}{4} \left( -\pi + 2 \tan^{-1} \left( \frac{\pi}{\sqrt{2}} \right) \right)$$

$$= \frac{\pi}{2} + \frac{\pi}{4} = \frac{3\pi}{4}$$

$$= 2.36$$

# 201ce Maths Paper 1

**Sean McArdle**

## 201ce Maths Paper 1:

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text o text c see table below notice that the starting concentrations of ce no and ce h

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web determination of a rate law and activation energy in this experiment glass beakers pipettes and a hotplate were utilized to set up the procedures a solution of deionized water a buffer potassium iodide ki sodium thiosulfate  $\text{Na}_2\text{S}_2\text{O}_3$  starch and hydrogen peroxide  $\text{H}_2\text{O}_2$  were used in order to calculate the rate of reaction

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