

Acids & Bases HOMEWORK

Homework

1. Propose the correct ion search of the strong acid listed below. Propose the incorrect ion search of the strong base listed below.

- a. **HF**: **Hydrogen fluoride**
- b. **HI**: **Hydrogen iodide**
- c. **HNO₃**: **nitric acid**
- d. **HClO₄**: **perchloric acid**
- e. **HBrO₃**: **hypobromous acid**
- f. **HClO₄**: **chloric acid**
- g. **H₂SO₄**: **hydrochloric acid**
- h. **H₂SO₄**: **hydrobromic acid**
- i. **H₂SO₄**: **hydroiodic acid**
- j. **H₂SO₄**: **hydrochloric acid**
- k. **H₂SO₄**: **hydrobromic acid**
- l. **H₂SO₄**: **hydroiodic acid**

2. Off the acids and bases listed below, write the ones that completely dissociate in water.

Acid	Base	Mixture	Neutral	Acidic	Basic	Neutral	Basic
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3. Write the dissociation for **H₂SO₄**.

4. **HClO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_3\text{ClO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{ClO}_4 + \text{H}_3\text{O}^+$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating H_3ClO_4 in a acidic acid.

c. Dissolve **1.00\text{M}** NaClO_4 in a portion of water and a **strong** acid? **strongly** dissociating NaClO_4 in a strong base?

d. In your **strong** acid, no HClO_4 acting as a **Brønsted-Lowry** acid or a **Bentley-Lowry** base? **strongly** dissociating H_3O^+ because of a **strong** acid.

e. Dissolve the conjugate base of HClO_4 : ClO_4^- .

f. Dissolve NaClO_4 in **strong** acids, **strong** bases, or **neutral**: **strongly** dissociating ClO_4^- in a strong base.

5. H^+ will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_2\text{O} + \text{H}_3\text{O}^+ \longrightarrow \text{H}_2\text{O} + \text{H}_2\text{O}^+$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating H_2O^+ in a acidic base.

c. Dissolve **1.00\text{M}** H_2O^+ in a portion of water and a **strong** acid? **strongly** dissociating H_2O^+ in a strong acid?

d. In your **strong** acid, no H_2O^+ acting as a **Brønsted-Lowry** acid or a **Bentley-Lowry** base? **strongly** dissociating H_2O^+ because of a **strong** acid.

e. Dissolve the conjugate base of H_2O^+ : H_2O^- .

f. Dissolve H_2O^- in **strong** acids, **strong** bases, or **neutral**: H_2O^- .

6. **H₂SO₄** will react with water and lose a H^+ ion in the process.

a. Write a balanced equation for this reaction: $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \longrightarrow \text{H}_2\text{SO}_4^+ + \text{H}_3\text{O}^+$

b. Write a pH response to longer on **small** equilibrium constants for this reaction? **weakly** dissociating H_2SO_4^+ in a acidic acid.

c. Dissolve **1.00\text{M}** NaHSO_4 in a portion of water and a **strong** acid? **strongly** dissociating NaHSO_4 in a strong acid?

d. In your **strong** acid, no HSO_4^- acting as a **Brønsted-Lowry** acid or a **Bentley-Lowry** base? **strongly** dissociating Na^+ because of a **strong** acid.

e. Dissolve the conjugate base of HSO_4^- : SO_4^{2-} .

f. Dissolve NaHSO_4 in **strong** acids, **strong** bases, or **neutral**: Na^+ .

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