

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from  $\text{pH}$  and  $\text{pOH}$ , and vice versa from  $\text{pH}$  and  $\text{pOH}$ .
3. Calculate the  $\text{pH}$  of a weak acid.

## Vocabulary

- |                   |                               |                  |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. $\text{pH}$                | 6. $\text{pOH}$  |

## Key Equation

- |                                     |                                  |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

## Part A Completion

Read the statement and fill in the blank with the word or phrase that will complete the statement.

1. When a substance ionizes, it releases hydrogen ions. The concentration of  $\text{H}^+$  in the autoionization of water is  $1.0 \times 10^{-7}$  moles per liter at 25°C and is equal to  $1.0 \times 10^{-7}$  moles per liter.
2. The  $\text{pH}$  of a solution is the negative logarithm of the concentration of  $\text{H}^+$  in the solution. The  $\text{pH}$  of a solution is  $-\log [\text{H}^+]$ .
3. The  $\text{pH}$  of a solution is the negative logarithm of the concentration of  $\text{H}^+$  in the solution. The  $\text{pH}$  of a solution is  $-\log [\text{H}^+]$ .
4. The  $\text{pH}$  of a solution is the negative logarithm of the concentration of  $\text{H}^+$  in the solution. The  $\text{pH}$  of a solution is  $-\log [\text{H}^+]$ .
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6. The  $\text{pH}$  of a solution is the negative logarithm of the concentration of  $\text{H}^+$  in the solution. The  $\text{pH}$  of a solution is  $-\log [\text{H}^+]$ .
7. The  $\text{pH}$  of a solution is the negative logarithm of the concentration of  $\text{H}^+$  in the solution. The  $\text{pH}$  of a solution is  $-\log [\text{H}^+]$ .
8. The  $\text{pH}$  of a solution is the negative logarithm of the concentration of  $\text{H}^+$  in the solution. The  $\text{pH}$  of a solution is  $-\log [\text{H}^+]$ .
9. The  $\text{pH}$  of a solution is the negative logarithm of the concentration of  $\text{H}^+$  in the solution. The  $\text{pH}$  of a solution is  $-\log [\text{H}^+]$ .
10. The  $\text{pH}$  of a solution is the negative logarithm of the concentration of  $\text{H}^+$  in the solution. The  $\text{pH}$  of a solution is  $-\log [\text{H}^+]$ .

1.  $1.0 \times 10^{-7}$ 2.  $1.0 \times 10^{-7}$ 3.  $1.0 \times 10^{-7}$ 4.  $1.0 \times 10^{-7}$ 5.  $1.0 \times 10^{-7}$ 6.  $1.0 \times 10^{-7}$ 7.  $1.0 \times 10^{-7}$ 8.  $1.0 \times 10^{-7}$ 9.  $1.0 \times 10^{-7}$ 10.  $1.0 \times 10^{-7}$ 11.  $1.0 \times 10^{-7}$ 12.  $1.0 \times 10^{-7}$

# 19 2 Hydrogen Ions And Acidity Section Review Answers

**Robert W. Avakian**



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