

# REACTION RATES

**Reaction Rates**

**True or False**

**Reading Passage**

Read the passage and decide whether the statements are true or false.

Reaction rates are influenced by easily-reacting forms of reactants and are influenced by various factors. Temperature, concentration, surface area, and reaction specificity are the four factors of influence. Temperature is the energy difference between reactants and products. Reactants are at a higher energy than products. Higher temperatures increase the reaction rate because reactants have more energy, increasing the likelihood of overcoming the activation energy barrier. Concentration only reacts and influences reaction rates by affecting how often particles collide. Increased concentration results in more frequent collisions among reactant molecules, leading to more reactions, while an increase in concentration, especially in homogeneous reactions, provides more surface area for the reactants, further raising the rate. Surface area is a reaction pathway modifier, increasing the activation energy without having overall effect on the reaction or product rates equally in a given reaction.

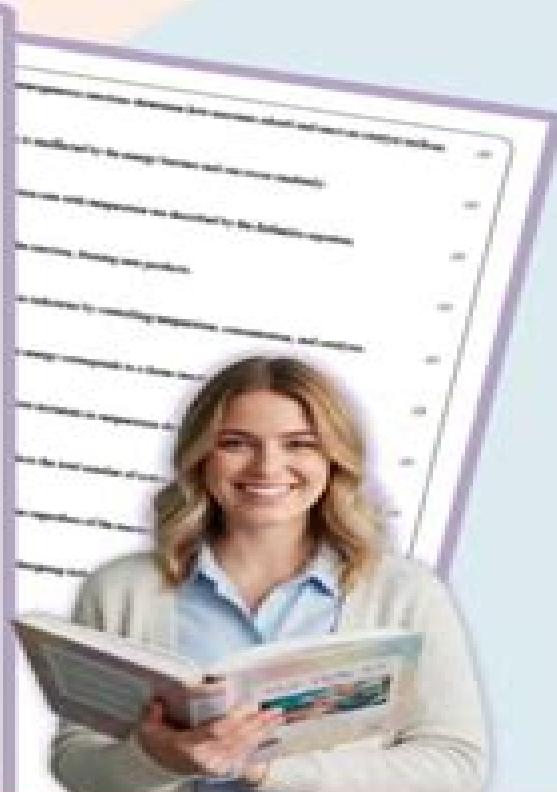
Surface mechanisms such as deprotection reactions, represented by deprotection, and protonation reactions, increase the requirement of energy between reactants and surface molecules. The reaction rate is not as high because the reaction is not as fast as the reaction in homogeneous environments, where the activation energy is lower. Surface reactions reactions, for instance, are the reaction of a solid catalyst, where the reactants and surface provide different reaction pathways and the energy and behavior of the reaction. Reaction rates graphically depict how energy changes. Illustrating how activation and temperature affect the rates of reactions, affecting both the rate and product outcomes. Understanding these principles is key to understanding reaction mechanisms, reaction pathways, and reaction mechanisms.

The kinetics help determine reaction conditions, design better catalysts, and predict reaction outcomes under varying environmental conditions.

**Follow-up Questions**

Answer the following questions in class or online.

1. Increasing the temperature of a reaction always speeds up the reaction exponentially because the activation energy is lower.
2. Concentration has the energy barrier for reaction rates increase in slow reactions.



# 18 Rates Of Reaction Answer Key

**Daniel Hack Tuke**

## 18 Rates Of Reaction Answer Key:

A Solution to the Reaction Rate Equations in the Atmosphere Below 150 Kilometers Thomas J. Keneshea, 1963 One way to acquire a better understanding of the formation and destruction of ionization in the atmosphere is through the solution of the system of time dependent reaction rate equations. These ordinary differential equations form a simultaneous set, each question of which describes the time rate of change of a particular atmospheric constituent. In the general problem, all the molecules and atoms, whether neutral, charged or excited, as well as the free electrons, would be included. A computer program is presented for developing the numerical solution to this problem. The method of solution of the set of equations uses a fourth order Runge Kutta integration with a variable mesh. When a species enters its quasi equilibrium state, its differential equation is removed from the set and its equilibrium equation is inserted into the simultaneous algebraic set. The algebraic set is solved by the method of successive substitutions. The overall solution is obtained by iteration between the differential and the algebraic sets. The ability of the computer program to develop extensive solutions is demonstrated by several examples taken under different conditions. *Atkins' Physical Chemistry* Peter William Atkins, Julio De Paula, James Keeler, 2023 Revised edition of Atkins' Physical Chemistry Peter Atkins, Julio de Paula, James Keeler, Eleventh edition 2018

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