

# ANSWER KEY: Ch. 17 Acid-base and Solubility Equilibria

**Question 1** Which of the following acids is not a weak acid? *1. Which of the following acids is not a weak acid?*

**Answer** D. *1. Which of the following acids is not a weak acid?*

The strong, nonvolatile acids and most acids are listed in concordance with acid-base literature. Which one of the following statements is not true?

- ☐ A. The equivalence point and end point will always be exactly the same for any titration.
- ☐ B. As the concentration of the analyte or titrant in the solution being titrated is the same as the number of equivalents of the titrant added.
- ☐ C. The end point of a titration is often dependent on an indicator color change.
- ☐ D. The equivalence point of a titration can be found using a pH meter.

**Question 2** *2. How many milliliters of 0.10 M NaOH are needed to neutralize 10 grams of HCl dissolved in water? (HCl is a strong acid.)*

**Answer** B. *2. How many milliliters of 0.10 M NaOH are needed to neutralize 10 grams of HCl dissolved in water? (HCl is a strong acid.)*

How many milliliters of 0.10 M NaOH are needed to neutralize 10 grams of HCl dissolved in water? (HCl is a strong acid.)

- ☐ A. 1000 mL
- ☐ B. 1000 mL
- ☐ C. 250 mL
- ☐ D. 250 mL

**Question 3** *3. The concentration of OH<sup>-</sup> in a acid solution whose pH is 4.*

**Answer** B. *3. The concentration of OH<sup>-</sup> in a acid solution whose pH is 4.*

The concentration of OH<sup>-</sup> in a acid solution whose pH is 4.

- ☐ A. 1.0 × 10<sup>-4</sup> mol/L
- ☐ B. 1.0 × 10<sup>-10</sup> mol/L

# 17 Answer Key Solubility Equilibria

**John T. Moore, Richard H. Langley**



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