

Common Polyatomic Ions

Formula	Name	Formula	Name
NH_4^+	Ammonium	CrO_4^{2-}	Chromate
NH_3	Ammonia	$\text{Cr}_2\text{O}_7^{2-}$	Dichromate
$\text{C}_2\text{H}_3\text{O}_2^-$	Acetate	MnO_4^-	Permanganate
CH_3COO^-	Acetate	MnO_4^{2-}	Manganate
CN^-	Cyanide	NO_2^-	Nitrite
CO_3^{2-}	Carbonate	NO_3^-	Nitrate
HCO_3^-	Bicarbonate	OH^-	Hydroxide
$\text{C}_2\text{O}_4^{2-}$	Oxalate	PO_4^{3-}	Phosphate
ClO^-	Hypochlorite	SCN^-	Thiocyanate
ClO_2^-	Chlorite	$\text{Fe}(\text{CN})_6^{3-}$	Ferricyanide

ClO_3^-	Chlorate	SO_3^{2-}	Sulfite
ClO_4^-	Perchlorate	SO_4^{2-}	Sulfate
$\text{S}_2\text{O}_3^{2-}$	Thiosulfate	HSO_4^-	Hydrogen sulfate
BrO^-	Hypobromite	IO_3^-	Iodate
AsO_2^{3-}	Arsenite	SeO_4^{2-}	Selenate
BrO_3^-	Bromate	HSO_3^-	Hydrogen

17 Polyatomic Ions S Answer Key

Paul B. Kelter, Steven S. Zumdahl

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CHEMISTRY
SILBERBERG, 2003 **The Solid-vacuum Interface** G. A. Bootsma, J. W. Geus, 1975 [Chemistry, Student Study Guide](#)

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Holt Chemistry Ralph Thomas Myers,2004 **Chemistry** Martin Stuart Silberberg,1996 This general chemistry text centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives

Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

Chemistry John S. Phillips,Cheryl Wistrom,2000 **Introductory Chemistry** Nivaldo J. Tro,2006 **Standard Reference Data Publications, 1987-1989** Joan C. Sauerwein,1989 **Test Bank to Accompany Chemistry and the Living Organism** Bloomfield,Catherine E. MacGowan,1995-12-08 **Introductory Chemistry** Charles H. Corwin,2005 For one semester courses in Basic Chemistry Introduction to Chemistry and Preparatory Chemistry and the first term of Allied Health Chemistry This text is carefully crafted to help students learn chemical skills and concepts more effectively Corwin covers math and problem solving early in the text he builds student confidence and skills through innovative problem solving pedagogy and technology formulated to meet student needs

Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 **The Chemistry of Everything** Kimberley Waldron,2007 The Chemistry of Everything addresses the need to know basics of chemistry required to grasp everyday science issues Through innovative themes and creative applications it provides an engaging introduction to chemistry for nonscience majors Mixes basic chemical principles from physical inorganic organic analytical and biological specializations to support thematic coverage of topics such as diamonds groceries and drugs Extends readers vocabulary and knowledge of the scientific issues encountered in daily life Addresses issues of ethics and responsible use in contemporary science Captures the current

fascination with forensics through Chemistry at the Crime Scene boxed sections For those interested in basic chemistry

Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter, Steven S. Zumdahl, 2004-04 **INIS**

Atomindeks, 1986 **Chemical Principles, Properties, and Reactions** Kenneth L. Henold, Frank Walmsley, 1984

Chemistry Charles H. Corwin, 1994 The book focuses on the concepts of chemistry and the applications that maintain and generate motivation for the subject of chemistry

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