

16.4**CALCULATIONS INVOLVING
COLLIGATIVE PROPERTIES****Section Review****Objectives**

- Calculate the molality and mole fraction of a solution
- Describe how the freezing-point depression and boiling-point elevation are related to molality

Vocabulary

- molality (m)
- mole fraction
- molal freezing-point depression constant (K_f)
- molal boiling-point elevation constant (K_b)

Key Equations

- Molality = $\frac{\text{moles of solute}}{\text{kilogram of solvent}}$
- mole fractions: $X_A = \frac{n_A}{n_A + n_B}$ $X_B = \frac{n_B}{n_A + n_B}$
 where n_A = moles of solute
 n_B = moles of solvent
- $\Delta T_b = K_b \times m$
- $\Delta T_f = K_f \times m$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Molality is an expression of concentration involving the ratio of 1 particles to 2 particles. Molality is expressed as moles of solute per 3 of solvent.

Another expression of concentration is 4, in which concentrations are expressed as the ratio of moles of solute to the total number of moles of solvent and solute. Each solvent has a characteristic 5 elevation constant and molal freezing-point 6 constant. The elevation in boiling point of a solution can be calculated by multiplying the 7 concentration of the solution by the boiling-point 8 constant of the solvent.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____

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Edward Mann Langley



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