

Semester 2 Final Exam Review – Answer Key

1. Mendeleev
2. Electrons in the outer-most energy shell of an atom
3. F, O, N
4. -ide
5. Atoms with a full valence shell
6. Molarity (M) = moles per liter = moles/liters
7. Basic, alkaline
8. Group or Family
9. $6.02 \times 10^{23} \times 5 = 3.01 \times 10^{24}$
10. Fluid, compressible, take the shape and volume of the container, molecules are far apart (low density)
11. Period
12. Acidic
13. High boiling point, low vapor pressure
14. A homogeneous mixture of only one phase; solute dissolved in a solvent
15. 22.4L
16. Endothermic
17. $Zn(NO_3)_2$
18. H, N, O, F, Cl, Br, I
19. Nature of the reactants, temperature, surface area, concentration, catalyst
20. Calorie = energy required to raise 1g of water 1 degree Celsius
21. Volume is directly proportional to temperature @ a constant pressure ($V_1/T_1 = V_2/T_2$)
22. Iron(II) chloride
23. $Ba(NO_3)_2$
24. $PbCO_3$
25. Reactants react to form products
26. Synthesis, decomposition, combustion, single replacement, double replacement
27. Water and carbon dioxide
28. Right side, $-\Delta H$
29. The energy require to form an activated complex and have a reaction proceed to form products
30. Exothermic, $\Delta H = -585\text{kJ}$
 - Convert mass to moles: $34\text{g} + 17\text{g/mol} = 2 \text{ mol NH}_3$
 - Multiply by energy/mole ratio: $2 \text{ mol NH}_3 \times (-1170\text{kJ})/4 \text{ mol NH}_3 = -585\text{kJ}$
31. Concentration, temperature, pressure (gas, heterogeneous system)
32. A buffer solution is a solution of a conjugate acid-base pair. A buffer system prevents significant changes in pH when a small amount of acid or base is added to the solution. Your body has a bicarbonate buffer system in the blood which helps maintain your blood's pH.
33. Unsaturated = more room for solute @ temp; saturated = solute maxed out @ temp; supersaturated = more solute than maximum for given temp

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Table of Contents Nd Semester Final Exam Review Science Answer

1. Understanding the eBook Nd Semester Final Exam Review Science Answer
 - The Rise of Digital Reading Nd Semester Final Exam Review Science Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Nd Semester Final Exam Review Science Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nd Semester Final Exam Review Science Answer
 - User-Friendly Interface

4. Exploring eBook Recommendations from Nd Semester Final Exam Review Science Answer
 - Personalized Recommendations
 - Nd Semester Final Exam Review Science Answer User Reviews and Ratings
 - Nd Semester Final Exam Review Science Answer and Bestseller Lists
5. Accessing Nd Semester Final Exam Review Science Answer Free and Paid eBooks
 - Nd Semester Final Exam Review Science Answer Public Domain eBooks
 - Nd Semester Final Exam Review Science Answer eBook Subscription Services
 - Nd Semester Final Exam Review Science Answer Budget-Friendly Options
6. Navigating Nd Semester Final Exam Review Science Answer eBook Formats
 - ePUB, PDF, MOBI, and More
 - Nd Semester Final Exam Review Science Answer Compatibility with Devices
 - Nd Semester Final Exam Review Science Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nd Semester Final Exam Review Science Answer
 - Highlighting and Note-Taking Nd Semester Final Exam Review Science Answer
 - Interactive Elements Nd Semester Final Exam Review Science Answer
8. Staying Engaged with Nd Semester Final Exam Review Science Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nd Semester Final Exam Review Science Answer
9. Balancing eBooks and Physical Books Nd Semester Final Exam Review Science Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nd Semester Final Exam Review Science Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Nd Semester Final Exam Review Science Answer
 - Setting Reading Goals Nd Semester Final Exam Review Science Answer
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Nd Semester Final Exam Review Science Answer
 - Fact-Checking eBook Content of Nd Semester Final Exam Review Science Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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