

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the K_a and K_b of acids and bases and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid-base indicator |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) indicated that are represented in the chemical equations below. Complete the reaction(s).

1. Water autoionizes. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. The autoionization of water which produces an H^+ ion and an OH^- ion. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. The autoionization of water which produces an H^+ ion and an OH^- ion. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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20. The autoionization of water which produces an H^+ ion and an OH^- ion. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

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Chao Zhang



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