

## Chemistry 2<sup>nd</sup> Semester Final Exam Review

### Chemical Bonds

- Give a physical description of how the atoms and molecules are arranged in solids, liquids, and gases.  
*A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces between the molecules (intermolecular forces) are stronger so they are held in place.*
- Use the information from the previous question to describe the shape and volume for each phase (state) of matter.  
*A: Solids have a repeating geometric pattern (lattice structure) in which the particles vibrate in place. They have a definite shape and a definite volume.  
Particles within a liquid slide past each and are close, but are not held in place. They have a definite volume, but no definite shape.  
Gas particles move chaotically bouncing off one another. They have neither a definite shape nor a definite volume.*
- List the three phases of matter in order of increasing intermolecular attractions.  
*A: Gas, liquid, solid.*
- Why do the atoms and molecules in liquids move in a random pattern relative to one another instead of being in a solid form?  
*A: Their kinetic energy (temperature) is high enough to overcome the attractive forces that might otherwise hold them in a repeating pattern. The attractive forces aren't strong enough to hold them together at that temperature.*

### Conservation of Matter and Stoichiometry

- The atomic mass unit and the mole are defined using what isotope?  
*A: Carbon-12*
- Balance the following equations:
  - $\text{Al}_2\text{O}_3 + \text{Cl}_2 + \text{C} \rightarrow \text{AlCl}_3 + \text{CO}$   
*A:  $\text{Al}_2\text{O}_3 + 3\text{Cl}_2 + 3\text{C} \rightarrow 2\text{AlCl}_3 + 3\text{CO}$*
  - $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$   
*A:  $2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 8\text{ CO}_2 + 6\text{H}_2\text{O}$*
- What isotope is used to define one mole?  
*A: Carbon-12*
- How many particles (atoms or molecules) are in one mole?  
*A:  $6.02 \times 10^{23}$*
- Find the molar mass for the following chemical formulas:
  - $\text{ZnF}_2$   
*A: 103.377 g/mol*
  - $\text{Al}_2(\text{SO}_4)_3$   
*A: 342.151 g/mol*
- How many moles are in 4.37 kg of NaOH?  
*A: 109.3 mol*
- How many atoms are in 2075 g of He?  
*A:  $3.12 \times 10^{26}$  atoms*
- What volume will 92.31 g of CO<sub>2</sub> occupy at STP?  
*A: 46.985 L*
- How many grams of sodium hydroxide are needed to react with 6.23 g of barium bromide?  
$$\text{NaOH} + \text{BaBr}_2 \rightarrow \text{NaBr} + \text{Ba(OH)}_2$$
  
*A: 1.68 g NaOH*
- In the reaction,  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ , if 100.0 g of magnesium reacts with 50.0 g of oxygen, what mass of product is produced?  
*A: 126 g MgO*

### Gases and Their Properties

- What is the kinetic molecular theory (KMT)?  
*A: The Kinetic Molecular Theory is used to explain the behavior of gases and is based upon the following postulates:
  - Gas particles are composed of a smoky particles that behave like hard spherical objects in a state of constant, random motion.
  - These particles move in a straight line until they collide with another particle or the walls of the container.
  - These particles are much smaller than the distance between particles, therefore the volume of a gas is mostly empty space and the volume of the gas molecule themselves is negligible.
  - There is no force of attraction between gas particles or between the particles and the walls of the container.
  - Collisions between gas particles or collisions with the walls of the container are elastic. That is, none of the energy of the gas particle is lost in a collision.*

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**M Tight**

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