

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the appropriate formula or name for each compound, and write its acid or base dissociation constant ( $K_a$  or  $K_b$ ) for each compound.

1.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

2.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

3.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

4.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

5.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

6.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

7.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

8.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

9.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

10.  $K_b = 1.8 \times 10^{-4}$  (Equation 19.2)

11.  $K_a = 1.8 \times 10^{-4}$  (Equation 19.1)

1.  $K_a = 1.8 \times 10^{-4}$
2.  $K_b = 1.8 \times 10^{-4}$
3.  $K_a = 1.8 \times 10^{-4}$
4.  $K_b = 1.8 \times 10^{-4}$
5.  $K_a = 1.8 \times 10^{-4}$
6.  $K_b = 1.8 \times 10^{-4}$
7.  $K_a = 1.8 \times 10^{-4}$
8.  $K_b = 1.8 \times 10^{-4}$
9.  $K_a = 1.8 \times 10^{-4}$
10.  $K_b = 1.8 \times 10^{-4}$
11.  $K_a = 1.8 \times 10^{-4}$

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