

**Concept-Development
Practice Page**

32-1

Coulomb's Law

1. The diagram is of a hydrogen atom.

- Label the proton in the nucleus with a + sign and the orbital electron with a - sign.
- The electrical interaction between the nucleus and the orbital electron is a force of _____ (attractive) _____ (repulsive).
- According to Coulomb's law,

$$F = k \frac{q_1 q_2}{r^2}$$

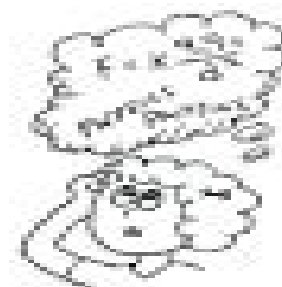
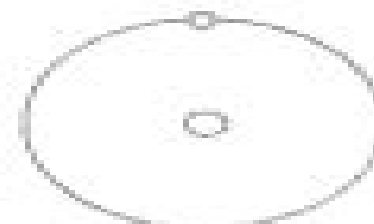
if the charge of either the nucleus or the orbital electron were greater, the force between the nucleus and the electron would be _____ (greater) _____ (less).

and if the distance between the nucleus and electron were greater the force would be _____ (greater) _____ (less).

If the distance between the nucleus and electron were doubled, the force would be _____ (1/4 as much) _____ (1/2 as much) _____ (two times as much) _____ (four times as much).

2. Consider the electric force between a pair of charged particles a certain distance apart. By Coulomb's law:

- If the charge on one of the particles is doubled, the force is _____ (quadrupled) _____ (halved) _____ (doubled) _____ (quadrupled).
- If instead the charge on both particles is doubled, the force is _____ (quadrupled) _____ (halved) _____ (doubled) _____ (quadrupled).
- If instead the distance between the particles is halved, the force is _____ (quadrupled) _____ (halved) _____ (doubled) _____ (quadrupled).
- If the distance is halved and the charge of both particles is doubled, the force is _____ times as great.



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Guillaume Favre



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Table of Contents 32 2 Concept Development

1. Understanding the eBook 32 2 Concept Development
 - The Rise of Digital Reading 32 2 Concept Development
 - Advantages of eBooks Over Traditional Books
2. Identifying 32 2 Concept Development
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 32 2 Concept Development
 - User-Friendly Interface
4. Exploring eBook Recommendations from 32 2 Concept Development
 - Personalized Recommendations
 - 32 2 Concept Development User Reviews and Ratings
 - 32 2 Concept Development and Bestseller Lists
5. Accessing 32 2 Concept Development Free and Paid eBooks
 - 32 2 Concept Development Public Domain eBooks
 - 32 2 Concept Development eBook Subscription Services
 - 32 2 Concept Development Budget-Friendly Options

6. Navigating 32 2 Concept Development eBook Formats
 - ePub, PDF, MOBI, and More
 - 32 2 Concept Development Compatibility with Devices
 - 32 2 Concept Development Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 32 2 Concept Development
 - Highlighting and Note-Taking 32 2 Concept Development
 - Interactive Elements 32 2 Concept Development
8. Staying Engaged with 32 2 Concept Development
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 32 2 Concept Development
9. Balancing eBooks and Physical Books 32 2 Concept Development
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 32 2 Concept Development
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 32 2 Concept Development
 - Setting Reading Goals 32 2 Concept Development
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 32 2 Concept Development
 - Fact-Checking eBook Content of 32 2 Concept Development
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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