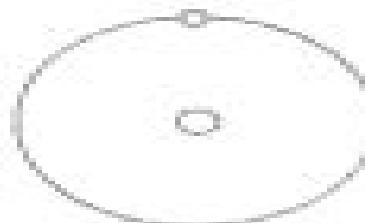


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Practice Page**
32-1
Consider this:

- The diagram is of a hydrogen atom.
- Label the parts in the nucleus with a + sign and the orbital electrons with a - sign.
- The electrical force between the nucleus and the orbital electrons is a force of _____ (attraction) (repulsion).
- According to Coulomb's law,

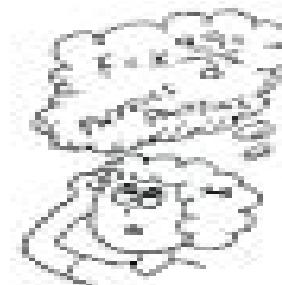
$$F = k \frac{q_1 q_2}{r^2}$$

If the charge of either the nucleus or the orbital electrons were greater, the force between the nucleus and the electron would be _____ (less) (more).
 If the distance between the nucleus and the orbital electrons were greater, the force would be _____ (less) (more).
 If the charges between the nucleus and electrons were doubled, the force would be _____ (4 times) (2 times) (less than or equal to) (more) as much.



- Consider the electric force between a pair of charged particles a certain distance apart. By Coulomb's law,

- If the charge on one of the particles is doubled, the force is _____ (quadrupled) (doubled) (unchanged) (reduced to one-half).
- If instead the charge on both particles is doubled, the force is _____ (quadrupled) (doubled) (unchanged) (reduced to one-half).
- If a constant distance between the particles is kept fixed, the force is _____ (quadrupled) (doubled) (unchanged) (reduced to one-half).
- If the distance is halved, and the charges of both particles is also halved, the force is _____ (none) (some) (more) (less) as great.



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Guillaume Favre

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