

QUESTION 3

3. 1.Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a $0,2 \text{ mol.dm}^{-3}$ ethanoic acid solution to 10 g of NaHCO_3 to start the reaction of the volcano.

The balanced equation for this reaction is:



- 3.1.1 Define the term limiting reagent. (2)
- 3.1.2 Determine the limiting reagent in this reaction. (6)
- 3.1.3 Calculate the mass of the other substance in excess. (3)
- 3.1.4 Calculate the volume of CO_2 produced at STP. (4)

3.2 A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in 250 cm^3 of water.

Calculate the concentration of the potassium hydroxide solution. (4)

3.3 25 cm^3 of a potassium hydroxide solution of concentration $0,25 \text{ mol.dm}^{-3}$ completely neutralises a dilute solution of sulphuric acid (H_2SO_4) in a flask.

The unbalanced equation below represents the reaction that takes place:



- 3.3.1 Balance the above equation. (2)
- 3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

[25]

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VM Jensen

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Table of Contents 2015 Msce Physical Science Practical Question

1. Understanding the eBook 2015 Msce Physical Science Practical Question
 - The Rise of Digital Reading 2015 Msce Physical Science Practical Question
 - Advantages of eBooks Over Traditional Books
2. Identifying 2015 Msce Physical Science Practical Question
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2015 Msce Physical Science Practical Question
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2015 Msce Physical Science Practical Question
 - Personalized Recommendations
 - 2015 Msce Physical Science Practical Question User Reviews and Ratings
 - 2015 Msce Physical Science Practical Question and Bestseller Lists
5. Accessing 2015 Msce Physical Science Practical Question Free and Paid eBooks
 - 2015 Msce Physical Science Practical Question Public Domain eBooks
 - 2015 Msce Physical Science Practical Question eBook Subscription Services
 - 2015 Msce Physical Science Practical Question Budget-Friendly Options
6. Navigating 2015 Msce Physical Science Practical Question eBook Formats

- ePub, PDF, MOBI, and More
 - 2015 Msce Physical Science Practical Question Compatibility with Devices
 - 2015 Msce Physical Science Practical Question Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2015 Msce Physical Science Practical Question
 - Highlighting and Note-Taking 2015 Msce Physical Science Practical Question
 - Interactive Elements 2015 Msce Physical Science Practical Question
 8. Staying Engaged with 2015 Msce Physical Science Practical Question
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2015 Msce Physical Science Practical Question
 9. Balancing eBooks and Physical Books 2015 Msce Physical Science Practical Question
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2015 Msce Physical Science Practical Question
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 2015 Msce Physical Science Practical Question
 - Setting Reading Goals 2015 Msce Physical Science Practical Question
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 2015 Msce Physical Science Practical Question
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 - Distinguishing Credible Sources
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 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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