

**4****Atomic Structure**  
Reviewsheet**A. Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

Atoms of each element are 1 from the atoms of all other elements. Dalton theorized that atoms are indivisible, but the discovery of 2 particles changed this theory. We now know that atoms are made up of electrons, which have a 3 charge; 4, which have a positive charge; and 5, which are neutral. The latter two particles are found in the 6 of the atom.

It was 7 who discovered the nucleus of the atom. The nucleus has a 8 charge and it occupies a very small volume of the atom. In contrast, the negatively charged 9 occupy most of the volume of the atom.

The number of 10 in the nucleus of the atom is the atomic 11 of that element. Because atoms are electrically neutral, the number of protons and 12 in an atom are equal. The sum of the 13 and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of 14 in the nucleus. Atoms that have the same number of protons but different mass numbers are called 15.

The 16 of an element is the weighted average of the masses of the isotopes of that element. Two isotopes of sulfur are  $^{32}\text{S}$  and  $^{34}\text{S}$ . An atom of the sulfur-32 isotope contains 17 protons and 18 neutrons. The sulfur-34 isotope has 19 protons and 20 neutrons.

1. different 4.1
2. subatomic 4.2
3. negative 4.2
4. protons 4.2
5. neutrons 4.2
6. nucleus 4.3
7. Rutherford 4.3
8. positive 4.3
9. electrons 4.3
10. protons (elec) 4.4
11. number 4.4
12. electrons 4.4
13. protons (elec) 4.6
14. neutrons 4.6
15. isotopes 4.6
16. atomic mass 4.7
17. 16 4.8
18. 16 4.8
19. 16 4.8
20. 18 4.8

## Atomic Structure Section Review Answers

**R Pring**

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## **Table of Contents Atomic Structure Section Review Answers**

1. Understanding the eBook Atomic Structure Section Review Answers
  - The Rise of Digital Reading Atomic Structure Section Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Section Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Atomic Structure Section Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Section Review Answers
  - Personalized Recommendations
  - Atomic Structure Section Review Answers User Reviews and Ratings
  - Atomic Structure Section Review Answers and Bestseller Lists
5. Accessing Atomic Structure Section Review Answers Free and Paid eBooks
  - Atomic Structure Section Review Answers Public Domain eBooks
  - Atomic Structure Section Review Answers eBook Subscription Services
  - Atomic Structure Section Review Answers Budget-Friendly Options
6. Navigating Atomic Structure Section Review Answers eBook Formats

- eBook, PDF, MOBI, and More
- Atomic Structure Section Review Answers Compatibility with Devices
- Atomic Structure Section Review Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Atomic Structure Section Review Answers
- Highlighting and Note-Taking Atomic Structure Section Review Answers
- Interactive Elements Atomic Structure Section Review Answers

8. Staying Engaged with Atomic Structure Section Review Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Atomic Structure Section Review Answers

9. Balancing eBooks and Physical Books Atomic Structure Section Review Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Atomic Structure Section Review Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Atomic Structure Section Review Answers

- Setting Reading Goals Atomic Structure Section Review Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Atomic Structure Section Review Answers

- Fact-Checking eBook Content of Atomic Structure Section Review Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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