

# ATOMIC STRUCTURE

Particle	Symbol	Charge	Relative charge	Mass	Relative mass w.r.t H	$\frac{e}{m}$ ratio
Electron	$e^-$	$-1.6 \times 10^{-19} \text{ C}$	-1	$9.1 \times 10^{-28} \text{ gram}$	$1/1837$	1837
Proton	$p^+$	$+1.6 \times 10^{-19} \text{ C}$	+1	$1.672 \times 10^{-24} \text{ g}$	1	1
Neutron	$n^0$	Zero	Zero	$1.672 \times 10^{-24} \text{ g}$	1	0
$\alpha$ -Particle	${}_2\text{He}^4$	$2 \times 1.6 \times 10^{-19} \text{ C}$	+2	$4 \times 1.672 \times 10^{-24} \text{ g}$	4	$1/2$

## Thomson's atomic model

(Plum pudding / Watermelon model)

- +ve charge  $\rightarrow$  uniformly distributed
- -ve charge  $\rightarrow$  embedded in cloud of the charge.
- Limitation  $\rightarrow$  doesn't explain stability

## Rutherford's experiment

(Discovery of nucleus)

Atom  $\rightarrow$  Nucleus (Proton)  
 $\rightarrow$  Extra nuclear part ( $e^-$ )

$$R_{\text{nucleus}} = K(A)^{1/3}$$

$$\left(\frac{R_{\text{nuc}}}{A}\right)_1 = \left(\frac{A_1}{A_2}\right)^{1/3}$$

$$V_{\text{nucleus}} = \frac{4}{3} \pi R^3$$

$$\frac{V_{\text{nuc1}}}{V_{\text{nuc2}}} = \frac{A_1}{A_2}$$

$$V_{\text{atom}} \approx 10^{15} \times V_{\text{nucleus}}$$

## Term associated with element

$Z$   $\swarrow$   $A \rightarrow$  Mass number ( $P+n$ )  
 $\searrow$  Atomic number ( $P$ )

- **Isotope**  $\rightarrow Z \rightarrow$  Same  $A \rightarrow$  different
- **Isobar**  $\rightarrow Z \rightarrow$  different  $A \rightarrow$  same
- **Isotone**  $\rightarrow A - Z = n \rightarrow$  same
- **Isotectonic**  $\rightarrow$  Same no. of  $e^-$  ( $\text{NO}^0, \text{CO}$ )
- **Isosteric**  $\rightarrow$  Same no. of atom and  $e^-$  ( $\text{CO}_2, \text{N}_2\text{O}$ )
- **Isotopomers**  $\rightarrow (n-p) \text{ or } (A-2Z) \rightarrow$  same
- **Paramagnetic**  $\rightarrow$  Having non zero unpaired  $e^-$
- **Diamagnetic**  $\rightarrow$  Having zero unpaired  $e^-$

## Dual nature of light

Wave nature

• Wave length ( $\lambda$ )  
 distance between crest or trough or length of one wave.

• Wave no.  $\boxed{\bar{\nu} = \frac{1}{\lambda}}$

• frequency  $\boxed{\nu = \frac{c}{\lambda}}$

$c \rightarrow 3 \times 10^8 \text{ m/s}$

$\nu \propto \frac{1}{\lambda} \rightarrow \boxed{\frac{\nu_1}{\nu_2} = \frac{\lambda_2}{\lambda_1}}$

Particle nature

• photon  $\rightarrow$  quanta of light.

$$\boxed{E = h\nu}$$

$h \rightarrow$  plank's constant  
 $h = 6.6 \times 10^{-34} \text{ joule sec}$

$$\boxed{E = n \times E_{\text{min}}}$$

$\downarrow$   
 integral multiple

$$\boxed{E = h\nu = \frac{hc}{\lambda}}$$

$$\boxed{\frac{E_1}{E_2} = \frac{\nu_1}{\nu_2}} \quad \boxed{\frac{E_1}{E_2} = \frac{\lambda_2}{\lambda_1}}$$

## Photo electric effect

- example of particle nature.
- Surface phenomena
- Single photon release single  $e^-$

$$\boxed{h\nu = \omega + KE} \quad \omega \rightarrow \text{work function}$$

$$\boxed{h\nu = h\nu_0 + \frac{1}{2} m_e v_e^2} \quad \nu \rightarrow \text{incident frequency of photon}$$

$$\boxed{\frac{hc}{\lambda} = \frac{hc}{\lambda_0} + \frac{1}{2} m_e v_e^2} \quad \nu_0 \rightarrow \text{Threshold frequency of metal.}$$

•  $\nu > \nu_0 \rightarrow e^-$  emitted with KE  
 $(\lambda < \lambda_0)$

•  $\nu = \nu_0 \rightarrow e^-$  emitted without KE  
 $(\lambda = \lambda_0)$

•  $\nu < \nu_0 \rightarrow e^-$  not emitted.  
 $(\lambda > \lambda_0)$

# Atomic Structure Quick Xword Answers

**Sylvie Doublie**



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