

ATOMIC STRUCTURE

Particle	Symbol	Charge	Relative charge	Mass	Relative mass w.r.t H	e/m ratio
Electron	e^-	$-1.6 \times 10^{-19} C$	-1	9.1×10^{-28} gram	$1/1837$	1837
Proton	p^+	$+1.6 \times 10^{-19} C$	+1	1.673×10^{-24} g	1	1
Neutron	n^0	Zero	Zero	1.678×10^{-24} g	1	0
α -Particle	${}^2He^4$	$1.6 \times 10^{-19} C$	+2	$4 \times 1.678 \times 10^{-24}$ g	4	$1/2$

Thomson's atomic model

(Plum pudding / Watermelon model)

- ve charge \rightarrow uniformly distributed
- ve charge \rightarrow embedded in cloud of the charge.
- Limitation \rightarrow can't explain stability

Rutherford's experiment

(Discovery of nucleus)

Atom \rightarrow Nucleus (Proton)
 \rightarrow Electrons (part (e^-))

$$\text{Nucleus} = K(A)^{1/3}$$

$$\frac{(\text{mass})_1}{(\text{mass})_2} = \left(\frac{A_1}{A_2}\right)^{1/3}$$

$$V_{\text{nucleus}} = \frac{4}{3} \pi R^3 \rho$$

$$\frac{V_{\text{nucleus}}}{V_{\text{atom}}} = \frac{A_1}{A_2}$$

$$V_{\text{atom}} \approx 10^{15} \times V_{\text{nucleus}}$$

Term associated with element

~~Z~~ \rightarrow Mass number ($P+N$)
 Z \rightarrow Atomic number (P)

- Isotope \rightarrow Z \rightarrow same A \rightarrow different
- Isobar \rightarrow Z \rightarrow different A \rightarrow same
- Isotone \rightarrow A-Z = 1 \rightarrow same
- Isoelectronic \rightarrow Same no. of e^- (NO³⁻, CO)
- Isotopes \rightarrow Same no. of atom and e⁻ (CO₂, NO₂)
- Isotrophies \rightarrow (1-P) or (A-2Z) \rightarrow same
- Paramagnetic \rightarrow Having non-zero unpaired e⁻
- Diamagnetic \rightarrow Having zero unpaired e⁻

Dual nature of light

Wave nature

- Wavelength (λ)
 distance between 2 crest or trough or length of one wave.
- Wave no. $\bar{v} = \frac{1}{\lambda}$

$$\text{frequency } v = \frac{c}{\lambda}$$

$$c \approx 3 \times 10^8 \text{ m/s}$$

$$\lambda \propto \frac{1}{\bar{v}} \rightarrow \frac{\bar{v}_1}{\bar{v}_2} = \frac{\lambda_2}{\lambda_1}$$

Particle nature

- photon \rightarrow quanta of light.

$$E = h\nu$$

$h \rightarrow$ plank's constant
 $h = 6.6 \times 10^{-34}$ joule sec

$$E = n \times E_{\text{min}}$$

\rightarrow integral multiple

$$E = h\nu = \frac{hc}{\lambda}$$

$$\frac{E_1}{E_2} = \frac{\nu_1}{\nu_2} \quad \frac{E_1}{E_2} = \frac{\lambda_2}{\lambda_1}$$

Photo electric effect

- example of particle nature.

- Surface phenomena

- Single photon release single e⁻

$$h\nu = \omega + KE \quad \omega \rightarrow \text{work function}$$

$$h\nu = h\nu_0 + \frac{1}{2}mv_e^2 \quad \nu \rightarrow \text{incident frequency of photon}$$

$$\frac{h\nu}{\lambda} = \frac{h\nu}{\lambda_0} + \frac{1}{2}mv_e^2 \quad \nu_0 \rightarrow \text{threshold frequency of metal.}$$

$\# \nu > \nu_0 \rightarrow e^-$ emitted with KE
 $(\lambda < \lambda_0)$

$\# \nu = \nu_0 \rightarrow e^-$ emitted without KE
 $(\lambda = \lambda_0)$

$\# \nu < \nu_0 \rightarrow e^-$ not emitted.
 $(\lambda > \lambda_0)$

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Sylvie Double

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