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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-H
Bond energy in kJ/mol	436	391

Use the bond energies in Table 3.1 and  $\Delta H$  to calculate the bond energy of an O-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.  

$$(400) + 1(945)$$

$$4(436) + 1(498)$$

$$3150 \text{ kJ}$$
- Calculate the energy released when bonds form in the products.  

$$\text{Bond broken} - \text{Bond formed} = -90$$

$$3150 - x = -90$$

$$x = 3240 + 90$$

$$3330 \text{ kJ}$$
- Calculate the energy of per mole of bond.  

$$6 \times \text{O-H} = 3330$$

$$\text{O-H} = \frac{3330}{6}$$

$$555 \text{ kJ/mol}$$

# 97accounting June 2013 Paper 42

**S Marginson**



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