



IGCSE A* MODEL ANSWER

0620 CHEMISTRY

May/June 2024

Paper 42



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QUESTION

(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.



Table 3.1 shows some bond energies.

Table 3.1

Bond	N≡N	H-H
Bond energy in kJ/mol	946	436

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an H-H bond in N₂H₄.

Use the following steps.

• Calculate the energy needed to break bonds in the reactants.

$$(946 \text{ kJ}) + 2(436 \text{ kJ})$$

$$1890 \text{ kJ} + 2(436 \text{ kJ})$$

1350

• Calculate the energy released when bonds form in the products.

$$\text{Reactants} - \text{Products} = -483.6 \text{ kJ}$$

$$1350 \text{ kJ} - x = -483.6 \text{ kJ}$$

$$x = 1833.6 \text{ kJ}$$

• Calculate the energy of an H-H bond.

$$\frac{1}{2} x \text{ (H-H)} = 917 \text{ kJ}$$

$$\frac{1}{2} x \text{ (H-H)} = \frac{917}{2} \text{ kJ}$$

$$458.5 \text{ kJ}$$

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