

# 11-2 Practice Problems

1. Determine the mass of a solution produced when 10.0 g of 1-butanol reacts with excess oxygen.

According to the following equation:  

$$C_4H_{10}O + 12O_2 \rightarrow 4CO_2 + 5H_2O$$

**Answer:** 26.9 g  $CO_2$  and 16.9 g  $H_2O$

2. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

According to the following equation:  

$$C_4H_{10}O + 12O_2 \rightarrow 4CO_2 + 5H_2O$$

3. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

According to the following equation:  

$$C_4H_{10}O + 12O_2 \rightarrow 4CO_2 + 5H_2O$$

4. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

According to the following equation:  

$$C_4H_{10}O + 12O_2 \rightarrow 4CO_2 + 5H_2O$$

5. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

6. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

7. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

According to the following equation:  

$$C_4H_{10}O + 12O_2 \rightarrow 4CO_2 + 5H_2O$$

8. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

9. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

10. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

11. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

12. Determine the mass of butane produced when 10.0 g of 1-butanol reacts with excess oxygen.

# 11 2 Practice Problems Answers Prentice Hall

**JA Banks**



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