

Stoichiometry & the mole notes

the mole concept

the mole (mol) is the SI unit for measuring the amount of a substance.

avogadro's number:

$$1 \text{ mol} = 6.022 \times 10^{23} \text{ particles}$$

(particles can be atoms, molecules, or formula units)

molar mass

the mass of one mole of a substance in grams.

found by adding the atomic masses from the periodic table

example

$$\text{H}_2\text{O} (2 \cdot 1.008 \text{ g/mol}) + (1 \cdot 16.00 \text{ g/mol}) = 18.016 \text{ g/mol}$$

Example Problems

1. how many atoms are in 3.5 moles of Fe?

$$3.5 \text{ mol Fe} \cdot \frac{6.022 \cdot 10^{23} \text{ atoms}}{1 \text{ mol}} = 2.11 \times 10^{24} \text{ atoms}$$

2. what is the mass of 2.00 moles of NaCl?

find molar mass:

$$\text{Na} : 22.99 \text{ g/mol}$$

$$\text{Cl} : 35.45 \text{ g/mol}$$

$$\text{Total} : 58.44 \text{ g/mol}$$

convert

$$2.00 \text{ mol} \cdot \frac{58.44 \text{ g}}{1 \text{ mol}} = 116.88 \text{ g}$$



Conversions Involving the mole

conversion

particles \leftrightarrow moles

moles \leftrightarrow mass

moles \leftrightarrow volume (L) (at STP)

factor to use

$$1 \text{ mol} = 6.022 \cdot 10^{23} \text{ particles}$$

$$1 \text{ mol} = \text{molar mass (g)}$$

$$1 \text{ mol} = 22.4 \text{ L of gas}$$

Stoichiometry: the study of quantitative relationships in chemical reactions

Balanced chemical equations show the mole ratio of reactants & products

mole ratios come from coefficients in a balanced equation.

Stoichiometry steps:

1. Balance the equation

2. convert given quantity into moles (if necessary)

3. Use mole ratio from balanced equation.

4. Convert moles back to desired unit (grams, liters, etc.)

Example:



question: how many grams of H_2O are produced from 5.00 moles of O_2 ?

Step 1: Use mole ratio

$$5.00 \text{ mol O}_2 \cdot \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} = 10.0 \text{ mol H}_2\text{O}$$

Step 2: Convert to grams

$$10.0 \text{ mol H}_2\text{O} \cdot \frac{18.016 \text{ g}}{1 \text{ mol}} = 180.16 \text{ g H}_2\text{O}$$

Limiting Reactant & Percent Yield

limiting reactant (LR): the reactant that runs out first, limiting product formation

excess reactant: the reactant that remains after the reaction is complete

Steps to identify LR

1. Convert reactants to moles

2. Use mole ratios to determine

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Table of Contents 9 Introduction To Stoichiometry Power Notes

1. Understanding the eBook 9 Introduction To Stoichiometry Power Notes
 - The Rise of Digital Reading 9 Introduction To Stoichiometry Power Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying 9 Introduction To Stoichiometry Power Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 9 Introduction To Stoichiometry Power Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from 9 Introduction To Stoichiometry Power Notes
 - Personalized Recommendations
 - 9 Introduction To Stoichiometry Power Notes User Reviews and Ratings
 - 9 Introduction To Stoichiometry Power Notes and Bestseller Lists
5. Accessing 9 Introduction To Stoichiometry Power Notes Free and Paid eBooks
 - 9 Introduction To Stoichiometry Power Notes Public Domain eBooks
 - 9 Introduction To Stoichiometry Power Notes eBook Subscription Services
 - 9 Introduction To Stoichiometry Power Notes Budget-Friendly Options
6. Navigating 9 Introduction To Stoichiometry Power Notes eBook Formats

- ePub, PDF, MOBI, and More
- 9 Introduction To Stoichiometry Power Notes Compatibility with Devices
- 9 Introduction To Stoichiometry Power Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 9 Introduction To Stoichiometry Power Notes
 - Highlighting and Note-Taking 9 Introduction To Stoichiometry Power Notes
 - Interactive Elements 9 Introduction To Stoichiometry Power Notes
- 8. Staying Engaged with 9 Introduction To Stoichiometry Power Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 9 Introduction To Stoichiometry Power Notes
- 9. Balancing eBooks and Physical Books 9 Introduction To Stoichiometry Power Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 9 Introduction To Stoichiometry Power Notes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 9 Introduction To Stoichiometry Power Notes
 - Setting Reading Goals 9 Introduction To Stoichiometry Power Notes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 9 Introduction To Stoichiometry Power Notes
 - Fact-Checking eBook Content of 9 Introduction To Stoichiometry Power Notes
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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