



LIMITING REAGENT AND PERCENT YIELD

Section 10.1

Limiting Reagents

DEFINITION

- limiting reagent: the reactant in a reaction mixture that is present in the smallest amount and thus limits the amount of product formed
- excess reagent: the reactant in a reaction mixture that is present in greater amounts than the limiting reagent

CHARACTERISTICS

- limiting reagent:
 - reacts first
 - reacts completely
 - remains unused
- excess reagent:
 - reacts later
 - reacts incompletely
 - remains in excess

Example 10.1

- $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$

Example 10.2

Given the following reaction, calculate the percent yield of CaCO_3 if 10.0 g of CaO and 10.0 g of CO_2 are used. Assume that the reaction is 100% complete.

$\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$

Given: $\text{CaO} = 10.0 \text{ g}$
 $\text{CO}_2 = 10.0 \text{ g}$

Find: CaCO_3

1. Calculate the number of moles of each reactant.

CaO moles = $\frac{10.0 \text{ g}}{56.08 \text{ g/mol}} = 0.178 \text{ mol}$

CO_2 moles = $\frac{10.0 \text{ g}}{44.01 \text{ g/mol}} = 0.227 \text{ mol}$

2. Calculate the number of moles of CaCO_3 produced.

From the balanced equation, 1 mol of CaO reacts with 1 mol of CO_2 to produce 1 mol of CaCO_3 .

Since CaO is the limiting reagent, the number of moles of CaCO_3 produced is 0.178 mol.

Percent Yield

- percent yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$
- theoretical yield = the amount of product that would be formed if the reaction were 100% complete
- percent yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$

Example 10.3

Given the following reaction, calculate the percent yield of CaCO_3 if 10.0 g of CaO and 10.0 g of CO_2 are used. Assume that the reaction is 100% complete.

- $\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$
- $\text{CaO} = 10.0 \text{ g}$
- $\text{CO}_2 = 10.0 \text{ g}$

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Conrad L. Stanitski

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