

ANSWERS TO STUDY GUIDE QUESTIONS
Romeo and Juliet
Study Guide – Act I

Prologue:

1. In what city does this play take place?

The play takes place in the Italian city of Verona.

2. Why are Romeo and Juliet called "Star-cross'd lovers?"

Romeo and Juliet are considered star-crossed lovers because fate has brought them together, yet their families are at war with one another and they will never be able to be together; therefore, fate will also tear them apart.

Act I, Scene I:

1. Who is fighting at the beginning of the first scene?

In the beginning of Act I Scene I, a fight occurs between the Capulets and the Montagues. The fight is between Sampson (Capulet), Gregory (Capulet), Abraham (Montague), and Balthazar (Montague). Later in the scene Benvolio (Montague) and Tybalt (Capulet) fight.

2. Who tries to break up the fighting?

Benvolio tries to break up the fighting but ends up joining the fight against Tybalt.

3. What threat does the Prince make to Lord Montague and Lord Capulet?

The Prince is tired of the fighting and threatens to execute anyone that becomes involved in a fight in the streets.

4. Benvolio and Montague describe the way Romeo has been acting. What do they have to say about him?

Benvolio and Montague talk about how depressed Romeo seems to be the last few days. Romeo is refusing to talk to anyone walking morosely through the gardens crying. Montague describes him as a flower that refuses to open up.

5. Why is Romeo so sad? Explain.

Romeo is in love with a woman named Rosaline. Rosaline has sworn that she will never love a man and wishes to join a convent (to become a nun). Romeo is sad that she does not love him back. Romeo feels alone and tells Benvolio that he will never be in love with another woman.

6. What is Benvolio's advice to Romeo?

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web bohr s model suggests that the atomic spectra of atoms are produced by electrons gaining energy from some source jumping up to a higher energy level then immediately dropping back to a lower energy level and emitting the energy difference between the two energy levels the existence of the atomic spectra is support for bohr s model of the atom

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web apr 5 2023 the aim of the experiment is to introduce the spectra emitted by some s block elements when their atoms

are excited by heating in a bunsen flame the experiment requires a hand spectroscope a bunsen burner flame test
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web the observable in a spectroscopy experiment is a spectrum which describes changes in the intensity and energy of light after interaction with the atom or molecule the goal of this experiment is to analyze the emission spectra of one electron and multi electron atoms and to use spectroscopy to identify unknown metals

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web this is a spectacular version of the flame tests experiment that can be used with chemists and non chemists alike it can be extended as an introduction to atomic spectra for post 16 students this experiment must be done as a demonstration it takes about ten minutes if all is prepared in advance

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web jan 1 2012 in experiment 19 the theory of atomic spectra we study the emission lines from the hydrogen atom in terms of the bohr theory of the atom the light coming from a hydrogen discharge tube is spread into its constituent wavelengths by a diffraction grating

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web experiment sweeping the detector arm through will now record a spectrum of the light from the discharge lamp try different aperture and slit sizes and adjusting the lens locations to record as many of the spectral peaks as possible dim spectral peaks will require careful tuning of the apertures to observe

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web the atomic spectra experiment is a complete solution for measuring the wavelengths of the discrete lines of the atomic spectra of various gases using a grating spectrophotometer providing educators worldwide with innovative solutions for teaching science

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web to observe and understand line emission spectra of atoms using gas discharge tubes to practice writing electron configurations for these and other elements electromagnetic radiation is energy in the form of waves waves are characterized by their wavelength λ and frequency ν

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web the field of atomic spectra provides in fact a perfect example of the way in which experiment and theory can react fruitfully on each other yielding advances in our understanding of the

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web the bohr model named after danish physicist niels bohr of an atom has a small positively charged central nucleus and electrons orbiting in at specific fixed distances from the nucleus

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web objectives to perform flame tests of metal cations in order to observe their featured colors to perform calculations up determine the frequency and energy in the emitted photons to tell these results the the types of electronic transitions occurring in are default

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web atomic emission spectra electrons in atoms normally occupy the lowest energy states possible such an atom is said to be in the ground state however electrons can be excited to high energy states when they absorb excess energy the excess energy can be provided by heat light or electrical discharge

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web atomic spectra experiment introduction in this experiment you will observe the spectra lines emitted by hydrogen gas and mercury gas contained in spectral tubes equipment optical bench meter stick diffraction grating lamp power supply hydrogen and mercury spectral tubes theory

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web specialists on hydrogen and hydrogen like systems presents the advances in theory and experiment that have occurred since that time behavior of electrons in atoms nov 30 2022 atomic spectra the interactions between atoms and electrons quantum theory of atomic structure

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more info the purpose of the lab was to determine the characteristics of visible light that was emitted by different elements we did this by conducting a flame test the purpose was

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web fingerprints of the elements atomic spectra as dr matilsky discussed in his video lecture atomic spectra occur due to the fact that orbital radii of electrons and hence their energies are quantized at specific levels determined by the atomic number number of protons and ionization state number of electrons in any given element

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web in this experiment the student measures the features of a simple spectrum that of sodium an empirical relation is deduced for the frequencies of the lines and the position of unobserved lines is predicted