

CHAPTER 8 REVIEW

Chemical Equations and Reactions**SECTION 1: Describing Chemical reactions**

1. Match the symbol on the left with its appropriate description on the right.

 Δ

(a) A precipitate forms.

 \downarrow

(b) A gas forms.

 \uparrow

(c) A reversible reaction occurs.

 (l)

(d) Heat is applied to the reactants.

 (aq)

(e) A chemical is dissolved in water.

 $\ddot{\cdot}$

(f) A chemical is in the liquid state.

2. Finish balancing the following equation:



3. In each of the following formulas, write the total number of atoms present.

 a. 4SO_2 b. 8O_2 c. $3\text{Al}_2(\text{SO}_4)_3$

4. Convert the following word equation into a balanced chemical equation: aluminum metal + copper(II) fluoride \rightarrow aluminum fluoride + copper metal



5. One way to test the salinity of a water sample is to add a few drops of silver nitrate solution with a known concentration. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, and sodium nitrate is left in solution. Translate these sentences into a balanced chemical equation.



2neutralization Reaction Section Review Answers

Aftab S. Hassan,James L. Flowers

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