

9.6 MC Questions



In a blast furnace, the reaction represented above occurs, producing $\text{Fe}_{(l)}$ from its ore, $\text{Fe}_2\text{O}_{3(l)}$.

The reaction is thermodynamically favorable and based on coupling the two reactions represented below.



Which of the following identifies a limitation in how the representations above describe a system of coupled reactions?

- A None of the equations are balanced.
- B The values of ΔS° for each reaction are not shown.
- C The values of ΔG° for each reaction are not shown.
- D The coupled reactions lack a common intermediate species.

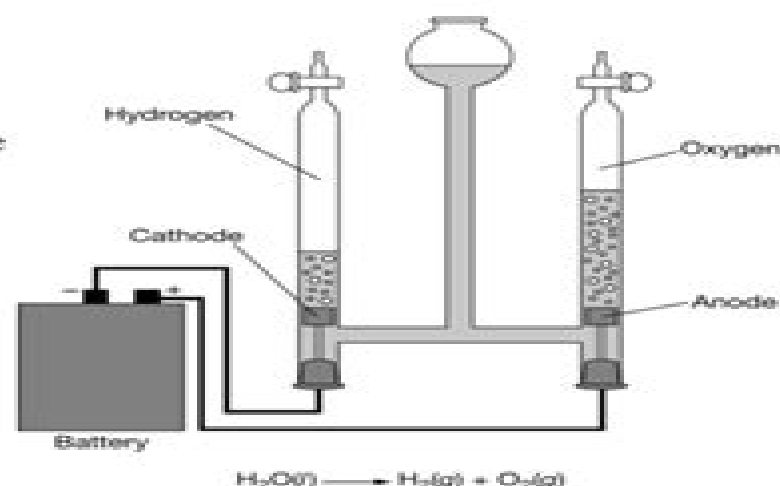
2. The reaction in which $\text{H}_2\text{O}_{(l)}$ is decomposed into $\text{H}_{2(g)}$ and $\text{O}_{2(g)}$ is thermodynamically unfavorable ($\Delta G^\circ > 0$).



However, an electrolytic cell, such as the one represented opposite, can be used to make the reaction occur.

Which of the following identifies a flaw in the representation?

- A Oxidation is occurring at the anode instead of at the cathode.
- B The equation for the reaction is not correctly balanced.
- C Electrical energy is needed for the reaction to proceed.
- D The relative volumes of collected gas shown in each tube are incorrect.



Ap Electrochemistry Problems

JG Myers



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Electrochemical Methods Allen J. Bard, Larry R. Faulkner, Henry S. White, 2022-05-31 The latest edition of a classic textbook in electrochemistry The third edition of *Electrochemical Methods* has been extensively revised to reflect the evolution of electrochemistry over the past two decades highlighting significant developments in the understanding of electrochemical phenomena and emerging experimental tools while extending the book's value as a general introduction to electrochemical methods This authoritative resource for new students and practitioners provides must have information crucial to a successful career in research The authors focus on methods that are extensively practiced and on phenomenological questions of current concern This latest edition of *Electrochemical Methods* contains numerous problems and chemical examples with illustrations that serve to illuminate the concepts contained within in a way that will assist both student and mid career practitioner Significant updates and new content in this third edition include An extensively revised introductory chapter on electrode processes designed for new readers coming into electrochemistry from diverse backgrounds New chapters on steady state voltammetry at ultramicroelectrodes inner sphere electrode reactions and electrocatalysis and single particle electrochemistry Extensive treatment of Marcus kinetics as applied to electrode reactions a more detailed introduction to migration and expanded coverage of electrochemical impedance spectroscopy The inclusion of Lab Notes in many chapters to help newcomers with the transition from concept to practice in the laboratory The new edition has been revised to address a broader audience of scientists and engineers designed to be accessible to readers with a basic foundation in university chemistry physics and mathematics It is a self contained volume developing all key ideas from the fundamental principles of chemistry and physics Perfect for senior undergraduate and graduate students taking courses in electrochemistry physical and analytical chemistry this is also an indispensable resource for researchers and practitioners working in fields including electrochemistry and electrochemical engineering energy storage and conversion analytical chemistry and sensors

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