

### AP Chemistry Chemical Kinetics Worksheet Answers

1.	Half-lives	Reactant Fraction Remaining
	1	1/2
	2	1/4
	3	1/8
	4	1/16
	5	1/32

2. 2<sup>nd</sup> Order; 1<sup>st</sup> Order

3. a. True    b. True  
c. False; As [Reactant] decreases the rate will decrease    d. True

4. a.  $-\frac{1}{2} \frac{\Delta[\text{O}_2]}{\Delta t} = \frac{1}{3} \frac{\Delta[\text{O}_2]}{\Delta t}$   
b.  $-\frac{1}{2} \frac{\Delta[\text{HOF}]}{\Delta t} = \frac{1}{2} \frac{\Delta[\text{HF}]}{\Delta t} = \frac{\Delta[\text{O}_2]}{\Delta t}$   
c.  $-\frac{\Delta[\text{N}_2]}{\Delta t} = -\frac{1}{3} \frac{\Delta[\text{H}_2]}{\Delta t} = \frac{1}{2} \frac{\Delta[\text{NH}_3]}{\Delta t}$

5. a. rate =  $k[\text{NO}_2]^2[\text{CO}]^0$  or  $k[\text{NO}_2]^2$   
b. rate = 1/4 the original value

6. The initial rate R =  $k[\text{B}]_0^2$

7. a.  $6.0 \times 10^{-3} / 1.2 \times 10^{-2} = (0.10/0.20)^x$   
 $X = 1$  so it is 1<sup>st</sup> order with respect to A  
 $1.1 \times 10^{-3} / 1.2 \times 10^{-2} = (0.15/0.05)^x$   
 $X = 2$  so it is 2<sup>nd</sup> order with respect to B  
**Rate Law =  $k[\text{A}][\text{B}]^2$**   
b. Overall it is 3<sup>rd</sup> order  
c. Probably not a single step since a 3-body collision would have to occur.

8. a. half life =  $\ln(1/2) = -5.0 \times 10^{-4} \text{s}^{-1} t$   
 $-\ln(0.5) / 5.0 \times 10^{-4} \text{s}^{-1} = \mathbf{1.4 \times 10^3 \text{s}}$   
b. Assign an original value of 1.00M  
 $\ln(0.1/1.0) = -5.0 \times 10^{-4} \text{s}^{-1} t$   
 $t = -\ln(0.1) / 5.0 \times 10^{-4} \text{s}^{-1} = \mathbf{4.6 \times 10^3 \text{s}}$

9. After 30.min (one half-life)  
 $P_{\text{HOF}} = \frac{1}{2} (100.\text{mmHg}) = 50.\text{mmHg}$   
 $P_{\text{HF}} = 50.\text{mmHg}$   
 $P_{\text{O}_2} = 25.\text{mmHg}$   
Calculating the rate constant k,  
 $\ln \left( \frac{[\text{HOF}]}{[\text{HOF}]_0} \right) = -kt$        $\ln(0.5) = -k (30.\text{min})$

# Ap Chemistry Chemical Kinetics Worksheet Answers

**Noel Roderick Waite**

## Ap Chemistry Chemical Kinetics Worksheet Answers:

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**CliffsNotes AP Chemistry** Bobrow Test Preparation Services, 2009-02-09

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