

AP Chem Unit 2

bonds

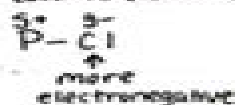
- based on electronegativity



* electronegativity = measure of how well atoms attract shared electrons

$X-X$
nonpolar
(same element)

$X-Y$
polar
(diff. element)



* greater diff. in electronegativity = more polar

* polar covalent bond = atom w/ greater electronegativity will have partial negative charge

- interstitial alloy: significant difference in atomic radius
smaller atoms fill in gaps



- substitutional alloy: 2 elements in which the atomic radius is similar



- metal + nonmetal = ionic

- nonmetal + nonmetal = covalent

- good conductor of electricity = metallic bonding

- solid doesn't conduct but as liquid does = ionic bonding

- solid that doesn't conduct & low melting point = covalent molecular

- solid that doesn't conduct & high melting point = covalent network

* metallic bonding = sea of mobile valence electrons
↳ good conductors



Bond Length and Bond Energy

$H-F$

$H-Cl$

$H-Br$

$H-I$

- shortest bond
- hard to break

- longest bond
- easy to break

$C-C$

$C=C$

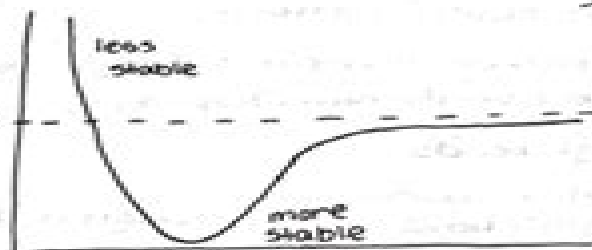
$C \equiv C$

- easy to break
- longest bond

- hard to break
- shortest bond

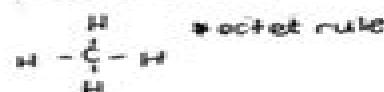
Periodic Table

as bond energy increases it becomes harder to break and bond length is shorter

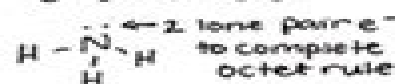


lewis dot diagrams

CH_4 = 8 valence e^-



NH_3 = 8 valence e^-



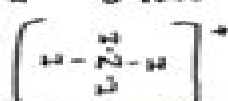
H_2O = 8 valence e^-



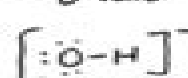
CH_2Cl_2 = 20 valence e^-



NH_4^+ = 8 valence e^-



OH^- = 8 valence e^-



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Magued Iskander



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