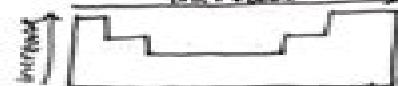


AP Chem Unit 2

bonds

- based on electronegativity increase



* electronegativity = measure of how well atoms attract shared electrons

- $X - X$ nonpolar (same element)
- $X - Y$ polar (diff. element)
- $H - F$
- $H - Cl$
- $H - Br$
- $H - I$

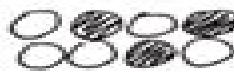
* polar covalent bond = atom with greater electronegativity will have partial negative charge

* more electronegative & greater diff. in electronegativity = more polar

- interstitial alloy: significant difference in atomic radius
- smaller atoms fill in gaps



- substitutional alloy: 2 elements in which the atomic radius is similar

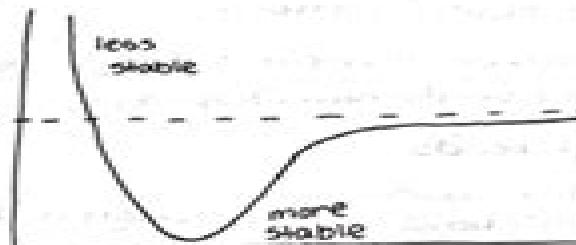


- metal + nonmetal = ionic
- nonmetal + nonmetal = covalent
- good conductor of electricity = metallic bonding
- solid doesn't conduct but as liquid does = ionic bonding
- solid that doesn't conduct & low melting point = covalent molecular
- solid that doesn't conduct & high melting point = covalent network
- * metallic bonding = sea of mobile valence electrons $[e^-]^{10-18}$
- * good conductors $[e^-]^{10-18}$

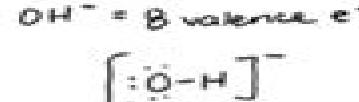
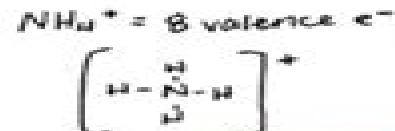
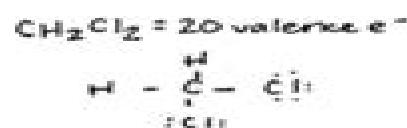
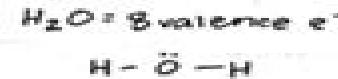
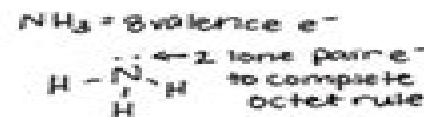
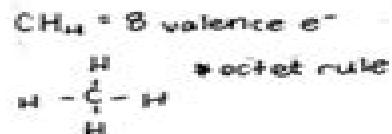
Bond Length and Bond Energy

- $H - F$ $H - Cl$ $H - Br$ $H - I$
- shortest bond
- hard to break
- $C - C$
- easy to break
- longest bond
- $C = C$
- hard to break
- shortest bond

as bond energy increases it becomes harder to break and bond length is shorter



Lewis dot diagrams



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Magued Iskander

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