

- Chapter Two: The Chemistry Of Life: Just When You Thought You Were Free, It's Back!
  - What is a chemical?
    - Chemical=any substance consisting of matter
      - matter=anything that takes up space and has mass
      - Organisms are composed of matter
  - Matter consists of chemical elements in pure form and in combinations called compounds.
    - Elements: substances that can't be broken down into other substances by chemical reactions
      - 92 naturally occurring elements, 118 total when including short-lived, man-made elements
    - Compound(s): substances consisting of two or more elements in a fixed ratio
      - A compound has characteristics that are different from those of the elements that compose it
      - emergent properties: become apparent as complexity increases
    - About 20-25% of 92 naturally occurring elements are essential elements necessary for life!
      - Carbon, hydrogen, oxygen, nitrogen make up 98% of living matter-think organic chemistry!
      - Trace elements: those required in small amounts, elements needed can vary from organism to organism - not every organism needs all, but every organism does need C,H,O,N-examples of trace elements include iron, zinc, magnesium
  - An element's properties depend on the structure of its atoms
    - atom=smallest unit of matter that still retains properties of an element
      - Made of subatomic particles (neutrons:  $\pm$  charge, protons: + charge, electrons: - charge)
      - Neutrons and protons form atomic nucleus, electrons form cloud around nucleus
    - Atomic number and atomic mass
      - Atoms of different elements differ in number of subatomic particles
      - Atomic number=number of protons
        - Determine identity of the atom
      - Mass number=protons+neutrons
        - Electrons weigh almost nothing! Atomic mass can be estimated using just protons and neutrons
  - All atoms have energy
    - Energy is the capacity to cause change
      - Potential energy: energy that matter has due to its position or structure (like a ball sitting on a high shelf)
      - An atom's electrons vary in the amount of energy they possess - takes more energy to move an electron away from the nucleus, so electrons further away have more potential energy
  - The chemical behavior of an atom is determined by the distribution of electrons in electron shells and from there how much space is available for bonding

# Ap Bio Chapter 1 outline

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