

- Chapter Two: The Chemistry Of Life: Just When You Thought You Were Free, It's Back!
 - What is a chemical?
 - Chemical=any substance consisting of matter
 - matter=anything that takes up space and has mass
 - Organisms are composed of matter
 - Matter consists of chemical elements in pure form and in combinations called compounds
 - Elements: substances that can't be broken down into other substances by chemical reactions
 - 92 naturally occurring elements, 118 total when including short-lived, man-made elements
 - Compounds: substances consisting of two or more elements in a fixed ratio
 - A compound has characteristics that are different from those of the elements that compose it
 - emergent properties: become apparent as complexity increases
 - About 20-25% of 92 naturally occurring elements are essential elements: necessary for life!
 - Carbon, hydrogen, oxygen, nitrogen make up 98% of living matter-think organic chemistry!
 - Trace elements: those required in small amounts, elements needed can vary from organism to organism - not every organism needs all, but every organism does need C,H,O,N-examples of trace elements include iron, zinc, magnesium
 - An element's properties depend on the structure of its atoms
 - atom=smallest unit of matter that still retains properties of an element
 - Made of subatomic particles (neutrons: \pm charge, protons: $+$ charge, electrons: $-$ charge)
 - Neutrons and protons form atomic nucleus, electrons form cloud around nucleus
 - Atomic number and atomic mass
 - Atoms of different elements differ in number of subatomic particles
 - Atomic number=number of protons
 - Determine identity of the atom
 - Mass number=protons+neutrons
 - Electrons weigh almost nothing! Atomic mass can be estimated using just protons and neutrons
 - All atoms have energy
 - Energy is the capacity to cause change
 - Potential energy: energy that matter has due to its position or structure (like a ball sitting on a high shelf)
 - An atom's electrons vary in the amount of energy they possess - takes more energy to move an electron away from the nucleus, so electrons further away have more potential energy
 - The chemical behavior of an atom is determined by the distribution of electrons in electron shells and from there how much space is available for bonding

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