

Antacid Analysis and Titration – Lab Report Assistant

Exercise 1: Back Titration of Antacid Neutralization

Data Table 1. Antacid Neutralization Data

Mass of Crushed Antacid (g)	
Concentration of HCl (M)	1M
Volume HCl (mL)	5 mL
Concentration of NaOH (M)	1M
Initial NaOH Volume (mL)	
Final NaOH Volume (mL)	
Total Volume of NaOH Used (mL)	

Data Table 2. Experimental Results

HCl available for neutralization (g):	0.3646 (g)
Moles of NaOH required to reach stoichiometric point:	0.0039 (mols)
HCl neutralized by antacid:	0.222 (g)
HCl neutralized per gram of antacid:	0.445 (g)

Questions:

- If an antacid tablet weighed 1.6 grams, how many moles of gastric acid (HCl) would it neutralize? Use the results obtained in Data Tables 1 and 2 to explain and quantify your answer.
- If you performed this experiment with a titrant of 0.5M NaOH, would you expect your results in Data Table 2 to change or stay the same? Explain your answer.
- The reaction that occurred when the antacid mixed with the HCl resulted in an additional product (besides a salt and water). Did you see evidence of this product? Describe the experimental evidence you witnessed that supports the formation of the additional product.



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BM King

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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques. The manual helps students understand the timing and situations for the various techniques. The Beran lab manual has long been a market leading lab manual for general chemistry. Each experiment is presented with concise objectives, a comprehensive list of techniques and detailed lab intros and step by step procedures.

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Introductory Chemistry Wendy Gloffke, Doris Kimbrough, 2002 This lab manual helps students develop data acquisition, organization and analysis skills while teaching basic techniques. Students construct their own data tables, answer conceptual questions and make predictions before performing experiments. They also have the opportunity to visualize and describe molecular level activity and explain the results.

Laboratory Manual for Fundamentals of General, Organic, and Biological Chemistry, Third Edition John R. Holum, Ruth C. Denison, Sandra L. Olmsted, 1986 This Laboratory Manual is designed to accompany the texts Fundamentals of General, Organic and Biological Chemistry 2nd Edition and Elements of General and Biological Chemistry 6th Edition by John R. Holum. It is also appropriate for any one year course treating a survey of chemistry at this level and for one term courses covering the whole spectrum of any part of it. The experiments have been used by students and have been frequently revised following student polls regarding clarity and interest and suggestions from instructors. The questions on the Report and Observation Sheets have again been adjusted in the light of student comments and more room for answers has been provided on many Report Sheets.

Analytical Chemistry for Technicians, Fourth Edition John Kenkel, 2013-08-13 Written as

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