

Unit Test 1: Matter, Chemical Trends, & Chemical Bonding**Multiple Choice (10 marks) (KU: 3, C: 3, I: 2)***Identify the choice that best completes the statement or answers the question. Place your answers in the key below.*

1. ____ 2. ____ 3. ____ 4. ____ 5. ____

6. ____ 7. ____ 8. ____ 9. ____ 10. ____

1. The element found in the periodic table in Group 6 and Period 5 is
 - a. tungsten, W
 - b. molybdenum, Mo
 - c. tantalum, Ta
 - d. tellurium, Te
2. Which of the following groups is known for having both metallic and non-metallic properties?
 - a. alkali metals
 - b. alkali earth metals
 - c. transition metals
 - d. metalloids
3. How many electrons does a strontium atom have in its outer shell?
 - a. 1
 - b. 2
 - c. 7
 - d. 8
4. Which particles are responsible for the similar chemical and physical properties of elements in the same group?
 - a. protons and neutrons
 - b. number of energy shells
 - c. valence electrons
 - d. isotopes
5. Which of the following chemical formulas most likely represents a pure substance containing ionic bonds?
 - a. Cl_2
 - b. SO_3
 - c. CO_2
 - d. BaO
6. Consider the equation $\text{X}_{(\text{g})} + e^- \rightarrow \text{X}^- + \text{energy}$. The "energy" term in the equation represents
 - a. electronegativity
 - b. electron affinity
 - c. ionization energy
 - d. nuclear charge
7. Elements A, B, C, and D (found in Groups 1-17) have atomic radii of $265 \text{ }\mu\text{m}$, $160 \text{ }\mu\text{m}$, $185 \text{ }\mu\text{m}$, and $175 \text{ }\mu\text{m}$, respectively. Which element will most likely have the lowest ionization energy?
 - a. A
 - b. B
 - c. C
 - d. D
8. List the following elements from smallest to largest atomic size:
 - I. barium
 - II. strontium
 - III. bismuth
 - IV. iodine
 - a. III, I, II, IV
 - b. IV, III, II, I
 - c. IV, III, III, I
 - d. IV, III, I, III
9. Which shape and bond angle is predicted by VSEPR theory for NF_3 ?
 - a. trigonal pyramidal, less than 109.5°
 - b. tetrahedral, 109.5°
 - c. trigonal planar, 109.5°
 - d. trigonal planar, 120°

Answers To Unit 5 Post Test Chemistry

Kendon Douglas Smith

Answers To Unit 5 Post Test Chemistry:

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