

# Chapter 10 Chemical Quantities Review

## Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- |                         |   |
|-------------------------|---|
| _____ mass              | _____ percent composition                                 |
| _____ empirical formula | _____ molar mass  |
| _____ molecular formula | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
1. \_\_\_\_\_ is a description of the relative amounts of each element in a compound.
2. \_\_\_\_\_ is the lowest whole-number ratio of the atoms of the elements in a compound.
3. \_\_\_\_\_ is the lowest structural formula and mass of any gas at STP.
4. \_\_\_\_\_ is the lowest structural formula of a compound.
5. \_\_\_\_\_ is the lowest structural formula of a compound.
6. \_\_\_\_\_ is the lowest structural formula of a compound.
7. \_\_\_\_\_ is the lowest structural formula of a compound.
8. \_\_\_\_\_ is the lowest structural formula of a compound.
9. \_\_\_\_\_ is the lowest structural formula of a compound.
10. \_\_\_\_\_ is the lowest structural formula of a compound.

## Part 2: Section Concept Review

### 10.1 The Mole: A Measurement

Complete the following statement by choosing the correct word or phrase. Write the correct word or phrase in the space provided.

1. \_\_\_\_\_ is the amount of substance that contains as many particles as there are atoms in 12 g of carbon-12.
2. \_\_\_\_\_ is the amount of substance that contains as many particles as there are atoms in 12 g of carbon-12.
3. \_\_\_\_\_ is the amount of substance that contains as many particles as there are atoms in 12 g of carbon-12.
4. \_\_\_\_\_ is the amount of substance that contains as many particles as there are atoms in 12 g of carbon-12.
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10. \_\_\_\_\_ is the amount of substance that contains as many particles as there are atoms in 12 g of carbon-12.

#### True or False?

1. \_\_\_\_\_ The mass of one mole of a substance is equal to its molar mass.

2. \_\_\_\_\_ The mass of one mole of a substance is equal to its molar mass.

3. \_\_\_\_\_ The mass of one mole of a substance is equal to its molar mass.

4. \_\_\_\_\_ The mass of one mole of a substance is equal to its molar mass.

5. \_\_\_\_\_ The mass of one mole of a substance is equal to its molar mass.

6. \_\_\_\_\_ The mass of one mole of a substance is equal to its molar mass.

#### Problems

# 10 Chemical Quantities Chapter Quiz Answer Key

**Barbara Ucko**



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**Chemistry: The Central Science** Theodore L. Brown,H. Eugene LeMay Jr.,Bruce E. Bursten,Catherine Murphy,Patrick Woodward,Steven Langford,Dalius Sagatys,Adrian George,2013-10-04

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*Study Guide to Accompany Basics for Chemistry* Martha Mackin,2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the

chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

**Student Guide For Living Chemistry** Barbara Ucko, 2012-12-02 Student Guide for Living Chemistry is a 23 chapter textbook guide that allows students to study and review on their own and test their understanding to help them prepare for examinations Every chapter begins with a list of objectives stating exactly the skills to develop in a particular unit Each objective corresponds to a section in the textbook Living Chemistry Three kinds of questions are provided for each objective to check the student's understanding namely short answer Study Questions multiple choice and fill in The answers for all questions are provided at the end of the chapter The opening chapters cover the SI units composition of matter chemical bonding compounds chemical change gases respiration and water The subsequent chapters deal with solutions acids bases salts nuclear and organic chemistry oxygen derivatives and hydrocarbons polymers and other organic derivatives This textbook also explores the chemistry of carbohydrates lipids proteins enzymes and energy and carbohydrate metabolism The remaining chapters discuss the chemistry of vitamins hormones body fluid drugs and poisons Undergraduate chemistry students will find this book invaluable

**CHEMISTRY** SILBERBERG, 2003 Foundations of Electronics Russell L. Meade, 1999 Learn electronics fundamentals for both DC and AC circuits from Ohms Law through series and parallel resonant circuits This highly acclaimed introduction to the world of electronics technology has been carefully updated to better provide technicians with a foundation in modern electronics needed to launch a career or pursue more advanced study Real world color codes and strategic highlighting are integrated with ample color charts photos schematics and diagrams for a solid understanding of circuit behavior that equips readers to progress to more complex topics with ease This edition features all new automated calculations for the formulas in the book on the accompanying CD as well as new information on admittance susceptance and more

**Brown's Introduction to Organic Chemistry** William H. Brown, Thomas Poon, 2017-06-28 Introduction to Organic Chemistry 6th Global Edition provides an introduction to organic chemistry for students who require the fundamentals of organic chemistry as a requirement for their major It is most suited for a one semester organic chemistry course In an attempt to highlight the relevance of the material to students the authors place a strong emphasis on showing the interrelationship between organic chemistry and other areas of science particularly the

biological and health sciences The text illustrates the use of organic chemistry as a tool in these sciences it also stresses the organic compounds both natural and synthetic that surround us in everyday life in pharmaceuticals plastics fibers agrochemicals surface coatings toiletry preparations and cosmetics food additives adhesives and elastomers Chemical Interactions ,2005 **Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed** Peter P. Berlow,1978 Merrill Chemistry Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 **A Practical Manual on Microbiologically Influenced Corrosion** Gregory Kobrin,1993 *Children's Books in Print* R R Bowker Publishing,Bowker,1999-12 **Living in the Environment** George Tyler Miller,2002 *Teaching Science for Understanding* James Joseph Gallagher,2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities **AQA GCSE Biology** Ann Fullick,2020-07-16 Specifically tailored for the 2016 AQA GCSE Science 9 1 specifications this third edition supports your students on their journey from Key Stage 3 and through to success in the new linear GCSE qualifications This series helps students and teachers to monitor progress while supporting the increased demand maths and new practical requirements *The Engineer* ,1885 **Scientific American** ,1872 **English Mechanic and Mirror of Science and Art** ,1886 **Science in Your World: Teacher edition** Jay K. Hackett,1991 **Introductory Chemistry** Nivaldo J. Tro,2006

## 10 Chemical Quantities Chapter Quiz Answer Key Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**10 Chemical Quantities Chapter Quiz Answer Key**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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