

**12****STOICHIOMETRY****SECTION 12.1 THE ARITHMETIC OF EQUATIONS** (pages 353–358)

This section explains how to calculate the amount of reactants required or product formed in a stoichiometric process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

**► Using Everyday Equations** (pages 353–355)

1. How can you determine the quantities of reactants and products in a chemical reaction?

You can use the balanced equation.

2. Quantity usually means the amount of a substance expressed in grams or moles.

3. A bookcase is to be built from 3 shelves (Sh), 2 side boards (Sb), 1 top (T), 1 base (B), and 4 legs (L). Write a "balanced equation" for the construction of this bookcase.

$3\text{Sh} + 2\text{Sb} + \text{T} + \text{B} + 4\text{L} = \text{Sh}_3\text{Sb}_2\text{TB}_1\text{L}_4$ .

**► Using Balanced Chemical Equations** (page 354)

4. Is the following sentence true or false? Stoichiometry is the calculation of quantities in chemical reactions. True

5. Calculations using balanced equations are called stoichiometric calculations.

**► Interpreting Chemical Equations** (pages 356–357)

6. From what elements is ammonia produced? How is it used?

Ammonia molecules are composed of nitrogen and hydrogen; it is used as a fertilizer.

7. Circle the letter of the term that tells what kind of information you CANNOT get from a chemical equation.

a. moles

d. volume

b. mass

e. number of particles

c. size of particles

## 10 3 Prentice Hall Chemistry Practice Problems

**Karin Nielsen-Saines**

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