

TSBIE - MODEL PAPER
CHEMISTRY - I

Time: 3 Hours

Max. Marks: 60

Each Question Carries TWO marks.

$2 \times 10 = 20$

SECTION - A

1. What is Lanthanide Contraction? Give one of its consequences.
2. Which of the two ions Ca^{2+} or Zn^{2+} is more stable and why?
3. What is Boltzman's constant? Give its value.
4. Calculate the weight of 0.1 mole of sodium carbonate.
5. What is homogenous equilibrium? Write two homogenous reactions.
6. Lithium salts are mostly hydrated. Why?
7. What happens when magnesium metal is burnt in air?
8. Diamond has high melting point – explain.
9. What is the effect of water on tin (Sn)?
10. Write the structure of a) Neopentane b) Trichloro ethanoic acid.

SECTION - B

Each Question Carries FOUR marks.

$6 \times 4 = 24$

11. Explain the differences between emission and absorption spectrum.
12. Balance the following reaction by ion electron method.
$$\text{Cr}_2\text{O}_7^{2-} + \text{SO}_3^{2-} \xrightarrow{\text{N}^+} \text{Cr}^{3+} + \text{SO}_4^{2-}$$
13. Write the postulates of Kinetic Molecular theory of gases.
14. Derive Ideal gas equation.
15. Chemical analysis of a carbon compound gave the following percentage composition by weight of the elements present, carbon = 10.06%, hydrogen = 0.84%, chlorine = 89.10%. Calculate the empirical formula of the compound.
16. State and explain Hess's law of constant heat summation.

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Table of Contents All Chemistry Inter Model Papers

1. Understanding the eBook All Chemistry Inter Model Papers
 - The Rise of Digital Reading All Chemistry Inter Model Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying All Chemistry Inter Model Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an All Chemistry Inter Model Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from All Chemistry Inter Model Papers
 - Personalized Recommendations
 - All Chemistry Inter Model Papers User Reviews and Ratings
 - All Chemistry Inter Model Papers and Bestseller Lists

5. Accessing All Chemistry Inter Model Papers Free and Paid eBooks
 - All Chemistry Inter Model Papers Public Domain eBooks
 - All Chemistry Inter Model Papers eBook Subscription Services
 - All Chemistry Inter Model Papers Budget-Friendly Options
6. Navigating All Chemistry Inter Model Papers eBook Formats
 - ePUB, PDF, MOBI, and More
 - All Chemistry Inter Model Papers Compatibility with Devices
 - All Chemistry Inter Model Papers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of All Chemistry Inter Model Papers
 - Highlighting and Note-Taking All Chemistry Inter Model Papers
 - Interactive Elements All Chemistry Inter Model Papers
8. Staying Engaged with All Chemistry Inter Model Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers All Chemistry Inter Model Papers
9. Balancing eBooks and Physical Books All Chemistry Inter Model Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection All Chemistry Inter Model Papers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine All Chemistry Inter Model Papers
 - Setting Reading Goals All Chemistry Inter Model Papers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of All Chemistry Inter Model Papers
 - Fact-Checking eBook Content of All Chemistry Inter Model Papers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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