

Counting by Measuring Mass

In chemistry, the particles (atoms, molecules, and ions) that react in chemical changes are extremely small, such that we need appropriate units to conveniently describe them. For example, a tablespoon of water contains approximately 6.02×10^{23} water molecules, give or take a few hundred quintillion. This is also known as 1 mole of water. 1 mole of any substance contains 6.022×10^{23} particles of that substance. If the substance is an element, then it contains 6.022×10^{23} atoms; if the substance is a molecular compound, then it contains 6.022×10^{23} molecules; if the substance is an ionic compound, then it contains 6.022×10^{23} formula units (the lowest ratio of ions that balances). This number is known as Avogadro's constant, named after the Italian scientist Amadeo Avogadro di Quaregna, Count of Turin (1776-1856), who helped distinguish molecules from atoms in gases. The Avogadro constant is defined as the amount of carbon-12 atoms in exactly 12 grams of carbon 12. Moles of a substance are a more convenient unit of count to work with than individual particles, similar to how it is more convenient to sell and purchase paper by the ream (300 sheets) instead by individual sheets of paper.

Since the particles are so small, directly measuring the mass of an individual atom is impossible, but measuring the mass by the mole is much more practical. The mass of an average oxygen atom is 2.66×10^{-23} grams, but the mass of 1 mole of oxygen atoms is 15.999 grams. This can be obtained from the periodic table, and it also represents the mass of an individual atom in atomic mass units (amu). So as obtained from the periodic table, the number 14.007 underneath the symbol for nitrogen (N) means that the mass of the average nitrogen atom is 14.007 amu, and the mass of 1 mole of nitrogen atoms is 14.007 grams, containing 6.022×10^{23} atoms of nitrogen. Therefore, using the atomic mass as a ratio, it is possible to determine the amount of nitrogen atoms in a given mass of nitrogen.

In this lab, you will determine the mass of several substances and use the data to count the number of atoms of each element in the substance.

EQ: How can mass be used to count the number of atoms in a substance?

Data Table

Part A	Al (s)	H ₂ O (l)	NaHCO ₃ (s)	C ₆ H ₁₂ O ₆ (s)
Mass (g)	4.123 grams	5.553 grams	4.616 grams	7.941 grams
Molar Mass (g/mol)	26.982 grams	18.015 grams	64.007 grams	180.16 grams
Moles of Each Substance	0.153 grams	0.308 grams	0.055 grams	0.044 grams
Number of Representative Particles	3.519×10^{23} atoms	7.084×10^{23} atoms	1.265×10^{23} atoms	1.014×10^{23} atoms
Moles of Each Element	0.153	H = .616 O = .308	Na = 0.055 H = 0.055 C = 0.055 O = 0.165	C = 0.264 H = 0.528 O = 0.264
Atoms of Each Element	3.519	H = 14.168 O = 7.084	Na = 1.265 H = 1.265 C = 1.265 O = 3.795	C = 6.084 H = 12.168 O = 6.084

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Table of Contents Answer Key Small Scale Lab

1. Understanding the eBook Answer Key Small Scale Lab
 - The Rise of Digital Reading Answer Key Small Scale Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Small Scale Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Small Scale Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Small Scale Lab
 - Personalized Recommendations
 - Answer Key Small Scale Lab User Reviews and Ratings

- Answer Key Small Scale Lab and Bestseller Lists
- 5. Accessing Answer Key Small Scale Lab Free and Paid eBooks
 - Answer Key Small Scale Lab Public Domain eBooks
 - Answer Key Small Scale Lab eBook Subscription Services
 - Answer Key Small Scale Lab Budget-Friendly Options
- 6. Navigating Answer Key Small Scale Lab eBook Formats
 - ePUB, PDF, MOBI, and More
 - Answer Key Small Scale Lab Compatibility with Devices
 - Answer Key Small Scale Lab Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Small Scale Lab
 - Highlighting and Note-Taking Answer Key Small Scale Lab
 - Interactive Elements Answer Key Small Scale Lab
- 8. Staying Engaged with Answer Key Small Scale Lab
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Small Scale Lab
- 9. Balancing eBooks and Physical Books Answer Key Small Scale Lab
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Small Scale Lab
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key Small Scale Lab
 - Setting Reading Goals Answer Key Small Scale Lab
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key Small Scale Lab
 - Fact-Checking eBook Content of Answer Key Small Scale Lab
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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