



## Unit: Chemical Bonding Answer Key

### "Covalent Bonding" - WS #3

Directions: Please answer each question using your notes from today's class. ☺

1. What is the difference between a molecular formula, structural formula and an electron dot formula? Give an example of each.

*A molecular formula gives the chemical make-up of one molecule. It does not show atom-to-atom connectivity. Example:  $C_2H_6$*

*A structural formula uses lines instead of dots to show shared pairs of electrons.*

*Example: H-H*

*An electron dot formula uses dots instead of lines to show shared pairs of electrons.*

*Example: H:H*

2. What is the difference between single, double, and triple covalent bonds? How many shared electrons are in each?

*The difference between the types of bonds is: a single bond is two shared electrons, a double is four shared electrons, and a triple is six shared electrons.*

3. Draw the electron dot and structural formulas for the following molecules.

Substance	Rough copy (work!)	Electron dot	Structural
a) $Cl_2$		<i>Note: The electron dot structures have dots in between each atom instead of lines.</i>	$:\ddot{Cl}—\ddot{Cl}:$
b) $H_2S$			$\begin{array}{c} \ddot{S} \\   \\ H \end{array}$
c) $BrI$			$\begin{array}{c} \ddot{I} \\   \\ Br \end{array}$

# Answer Chemical Bonding Ch 13

**Michal Rosen-Zvi**

## Answer Chemical Bonding Ch 13:

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*A Level Chemistry MCQ (Multiple Choice Questions)* Arshad Iqbal, 2019-06-18

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