Name Mary Date Class

TEACHING TRANSPARENCY WORKSHEET 31

The Activity Series

Use with Chapter 9, Section 9.2

1. For each of the following pairs of elements, underline the one that would replace the other element in a compound.
   a. calcium, tin
   b. bromine, fluorine
   c. aluminum, potassium
   d. zinc, sodium
   e. iron, copper
   f. iodine, chlorine
   g. silver, lead

2. For each of the following reactions, use the activity series to determine whether the reaction would take place or not. If no reaction takes place, write NR in the blank. If a reaction does take place, write the formulas for the products of the reaction. (Hint: If an active metal replaces the hydrogen in water, the hydroxide of the active metal forms.)
   a. Li(s) + $Fe(NO_3)_2$(aq) → $LiNO_3$ + Fe
   b. Au(s) + HCl(aq) → No Rxn.
   c. Cl_2(g) + KBr(aq) → KCl + Br_2
   d. Ca(s) + $Al(NO_3)_3$(aq) → No Rxn
   e. Ag(s) + HBr(aq) → No Rxn.
   f. Ni(s) + $SnCl_2$(aq) → $NiCl_2$ + Sn

3. Magnesium metal can be used to remove tarnish from silver items. Silver tarnish is the corrosion that occurs when silver metal reacts with substances in the environment, especially those containing sulfur. Why would magnesium remove tarnish from silver?
   My more reactive would replace Silver bounded to Sulfur (environmental reasons) leaving silver

4. Use the activity series for metals to explain why copper metal is used in plumbing where the water might contain compounds of many different metals.
   Cu not very reactive - won't react w/ metals and water

5. The last four metals in the activity series of metals are commonly referred to as the "coinage metals." Why would these metals be chosen over more active metals for use in coins? Why do you think some more active metals, such as zinc or nickel, are sometimes used in coins?
   lighter, more economical

Answer Key For Teaching Transparency

Minjie Lin

Answer Key For Teaching Transparency:

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