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## Reinforce and Review

### Lesson Worksheets

Copy and distribute the Lesson 12.3 worksheets in CK-12 Chemistry –Intermediate Workbook. Ask students to complete the worksheets to reinforce lesson content.

### Lesson Review Questions

Have students answer the Lesson Review Questions at the end of Lesson 12.3 in CK-12 Chemistry –Intermediate FlexBook® resource.

### Sample Answers

#### Practice Problems

1.

$$1.30 \text{ mol } C_3H_8 \times \frac{5 \text{ mol } O_2}{1 \text{ mol } C_3H_8} = 6.50 \text{ mol } O_2 \text{ needed}$$

Since the moles of  $O_2$  needed to fully react with all of the  $C_3H_8$  is more than the amount of  $O_2$  present (6.00 mol),  $O_2$  is the limiting reactant.

2.

$$8.00 \text{ g } Zn \times \frac{1 \text{ mol } Zn}{65.38 \text{ g } Zn} \times \frac{2 \text{ mol } HCl}{1 \text{ mol } Zn} \times \frac{36.46 \text{ g } HCl}{1 \text{ mol } HCl} = 8.92 \text{ g } HCl \text{ needed}$$

Since there is less HCl present (8.00 g) than what is needed to react with all of the Zn, the HCl is the limiting reactant. This problem can also be solved by calculating the moles of Zn and the moles of HCl originally present. Compare the number of moles, taking note of the 1:2 ratio.

$$\begin{aligned} 8.00 \text{ g } Zn \times \frac{1 \text{ mol } Zn}{65.38 \text{ g } Zn} &= 0.122 \text{ mol } Zn \\ 8.00 \text{ g } HCl \times \frac{1 \text{ mol } Zn}{36.46 \text{ g } HCl} &= 0.219 \text{ mol } HCl \end{aligned}$$

The moles of HCl is not double the moles of Zn, as required by the balanced equation. Therefore, HCl is the limiting reactant.

3.

- The balanced equation requires twice as many moles of Na as  $Cl_2$ . Sodium (Na) is the limiting reactant.
- The 3.40 moles of Na can be converted to moles of  $Cl_2$  that would react.

$$3.40 \text{ mol } Na \times \frac{1 \text{ mol } Cl_2}{2 \text{ mol } Na} = 1.70 \text{ mol } Cl_2 \text{ reacts}$$

The excess  $Cl_2$  is determined by subtraction.  $4.20 \text{ mol} - 1.70 \text{ mol} = 2.50 \text{ mol excess } Cl_2$

c.  $3.40 \text{ mol } Na \times \frac{2 \text{ mol } NaCl}{2 \text{ mol } Na} = 3.40 \text{ mol } NaCl$

- First, the limiting reactant must be determined.

$$35.0 \text{ g } N_2 \times \frac{1 \text{ mol } N_2}{28.02 \text{ g } N_2} \times \frac{3 \text{ mol } H_2}{1 \text{ mol } N_2} \times \frac{2.02 \text{ g } H_2}{1 \text{ mol } H_2} = 7.57 \text{ g } H_2 \text{ needed}$$

## 12 3 Review And Reinforce Answers

**Cynthia Y. Young**

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web basically khuswant singh wants to convey the irony of one s beliefs in the mark of vishnu in the story gunga ram is a devotee of vishnu vishnu is the preserver and protector of creation for

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web feb 15 2020 the mark of vishnu khushwant singh s tale of gunga ram highlights the perils of falling into the trap of blind faith and superstition khushwant singh published feb 16 2020 00 00 00 ist the indian cobra with the characteristic v mark on its hood image for representative purposes only courtesy biodiversity heritage library flickr

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web in the mark of vishnu by khushwant singh we have the theme of tradition modernity faith loyalty respect and change taken from his collected stories collection the story is narrated in the first person by a young unnamed boy and after reading the story the reader realises that singh may be exploring the theme of tradition

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web the theme of the short story the mark of vishnu by khushwant singh is superstition and ignorance the main character gunga ram is both superstitious and ignorant he believes that by leaving a saucer full of milk for the snake to drink he is protecting the household from snake bites

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web the main characters in the short story the mark of vishnu by khushwant singh are gunga ram and the four schoolboys gunga ram is a servant in the household where the schoolboys live the four brothers go to school and receive a formal education since gunga ram is a servant the schoolboys make fun of him for his lack of education and