

**12.3****LIMITING REAGENT AND PERCENT YIELD****Section Review****Objectives**

- Identify and use the limiting reagent in a reaction to calculate the maximum amount of product(s) produced and the amount of excess reagent
- Calculate theoretical yield, actual yield, or percent yield given the appropriate information

**Vocabulary**

- limiting reagent
- excess reagent
- theoretical yield
- actual yield
- percent yield

**Key Equations**

- $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

**Part A Completion**

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Whenever quantities of two or more reactants are given in a stoichiometric problem, you must identify the 1. This is the reagent that is completely 2 in the reaction. The amount of limiting reagent determines the amount of 3 that is formed.

When an equation is used to calculate the amount of product that will form during a reaction, the value obtained is the 4. This is the 5 amount of product that could be formed from a given amount of reactant. The amount of product that forms when the reaction is carried out in the laboratory is called the 6.

1. \_\_\_\_\_
2. \_\_\_\_\_
3. \_\_\_\_\_
4. \_\_\_\_\_
5. \_\_\_\_\_
6. \_\_\_\_\_

**Part B True-False**

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- \_\_\_\_\_ 7. Normally, the actual yield in a chemical reaction will be equal to or less than the theoretical yield.
- \_\_\_\_\_ 8. The actual yield of a chemical reaction can be calculated using mole ratios.

# 12 3 Limiting Reagent And Percent Yield Answer Key

**Morris Hein, Susan Arena**



### 12 3 Limiting Reagent And Percent Yield Answer Key:

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Holt Chemistry File ,1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket

**Practice Makes Perfect Chemistry Review and Workbook, Second Edition** Marian DeWane,Heather Hattori,2018-12-28 The Winning Equation for Success in Chemistry is Practice Practice Practice This book will help you apply concepts and see how chemistry topics are interconnected Inside are numerous lessons to help you better understand the subject These lessons are accompanied by dozens of exercises to practice what you ve learned along with a complete answer key to check your work Throughout this book you will learn the terms to help you understand chemistry and you will expand your knowledge of the subject through hundreds of sample questions and their solutions With the lessons in this book you will find it easier than ever to grasp chemistry concepts And with a variety of exercises for practice you will gain confidence using your growing chemistry skills in your classwork and on exams

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