

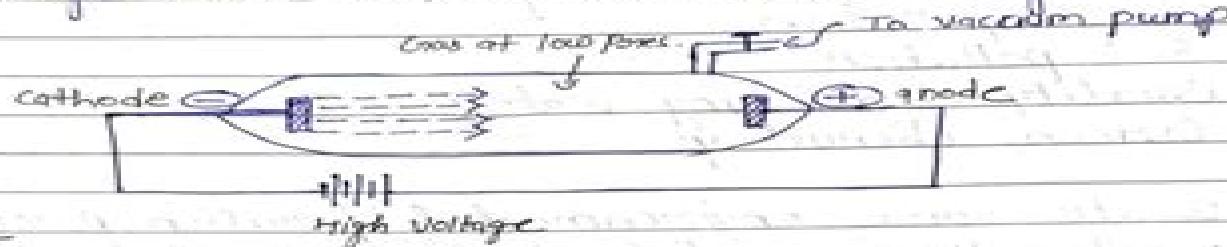
CHAPTER - 2

STRUCTURE OF ATOM

Discovery of electron (Cathode Rays)

Cathode rays \rightarrow William Crookes (1879)
Electron (e^-) \rightarrow J. J. Thomson. (-ve charge particle)

-ve charged rays which move from cathode to anode in discharge tube known as cathode rays.



Properties

- 1 Cathode rays travel in a straight line (Straight)
- 2 Cathode rays carry -ve charge (negative rays, deflected towards positive)
- 3 Cathode rays are deflected by magnetic field
- 4 They are made up of material particle (Rata, saddle, wool)
- 5 They produce heating effect & possess penetrating effect
- 6 They produce green fluorescence & X-rays (Other rays)
- 7 They also affect photographic plate.

Discovery of proton (Anode / canal rays)

Anode rays \rightarrow Goldstein (1886)

Proton (p^+) \rightarrow Goldstein & named by Rutherford (+ve charge particle)

+ve charged particle which move opposite the cathode in discharge tube known as anode rays.

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Table of Contents 11th Class Up Board Chemistry Notes

1. Understanding the eBook 11th Class Up Board Chemistry Notes
 - The Rise of Digital Reading 11th Class Up Board Chemistry Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying 11th Class Up Board Chemistry Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 11th Class Up Board Chemistry Notes
 - User-Friendly Interface

4. Exploring eBook Recommendations from 11th Class Up Board Chemistry Notes
 - Personalized Recommendations
 - 11th Class Up Board Chemistry Notes User Reviews and Ratings
 - 11th Class Up Board Chemistry Notes and Bestseller Lists
5. Accessing 11th Class Up Board Chemistry Notes Free and Paid eBooks
 - 11th Class Up Board Chemistry Notes Public Domain eBooks
 - 11th Class Up Board Chemistry Notes eBook Subscription Services
 - 11th Class Up Board Chemistry Notes Budget-Friendly Options
6. Navigating 11th Class Up Board Chemistry Notes eBook Formats
 - ePUB, PDF, MOBI, and More
 - 11th Class Up Board Chemistry Notes Compatibility with Devices
 - 11th Class Up Board Chemistry Notes Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 11th Class Up Board Chemistry Notes
 - Highlighting and Note-Taking 11th Class Up Board Chemistry Notes
 - Interactive Elements 11th Class Up Board Chemistry Notes
8. Staying Engaged with 11th Class Up Board Chemistry Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 11th Class Up Board Chemistry Notes
9. Balancing eBooks and Physical Books 11th Class Up Board Chemistry Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 11th Class Up Board Chemistry Notes
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 11th Class Up Board Chemistry Notes
 - Setting Reading Goals 11th Class Up Board Chemistry Notes
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 11th Class Up Board Chemistry Notes
 - Fact-Checking eBook Content of 11th Class Up Board Chemistry Notes
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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