

11-2 Practice Problems

1. Determine the mass of a solution produced when 10.0 g of 1-butanol reacts with excess acetic acid.

According to the following equation:
 $\text{C}_4\text{H}_9\text{OH} + \text{CH}_3\text{COOH} \rightarrow \text{C}_4\text{H}_9\text{OOCCH}_3 + \text{H}_2\text{O}$



2. Determine the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.



3. Determine the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.



4. Determine the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.



5. Determine the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.



6. Determine the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.



7. Determine the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.

According to the following equation:
 $\text{C}_4\text{H}_9\text{OH} + \text{CH}_3\text{COOH} \rightarrow \text{C}_4\text{H}_9\text{OOCCH}_3 + \text{H}_2\text{O}$

8. Determine the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.



9. Determine the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.

10. Find the mass of water produced when 10.0 g of 1-butanol reacts with excess acetic acid.

11. Determine the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.



12. Find the mass of butanol produced when 10.0 g of 1-butanol reacts with excess acetic acid.



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AN Whitehead



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