

# Chemistry Study Guide

# Periodic Table of the Elements

<small>1</small> <b>H</b> <small>1.008</small>																	<small>2</small> <b>He</b> <small>4.003</small>																															
<small>3</small> <b>Li</b> <small>6.941</small>	<small>4</small> <b>Be</b> <small>9.012</small>											<small>5</small> <b>B</b> <small>10.81</small>	<small>6</small> <b>C</b> <small>12.01</small>	<small>7</small> <b>N</b> <small>14.01</small>	<small>8</small> <b>O</b> <small>16.00</small>	<small>9</small> <b>F</b> <small>18.998</small>	<small>10</small> <b>Ne</b> <small>20.18</small>																															
<small>11</small> <b>Na</b> <small>22.99</small>	<small>12</small> <b>Mg</b> <small>24.31</small>											<small>13</small> <b>Al</b> <small>26.98</small>	<small>14</small> <b>Si</b> <small>28.09</small>	<small>15</small> <b>P</b> <small>30.97</small>	<small>16</small> <b>S</b> <small>32.07</small>	<small>17</small> <b>Cl</b> <small>35.45</small>	<small>18</small> <b>Ar</b> <small>39.95</small>																															
<small>19</small> <b>K</b> <small>39.10</small>	<small>20</small> <b>Ca</b> <small>40.08</small>	<small>21</small> <b>Sc</b> <small>44.96</small>	<small>22</small> <b>Ti</b> <small>47.88</small>	<small>23</small> <b>V</b> <small>50.94</small>	<small>24</small> <b>Cr</b> <small>52.00</small>	<small>25</small> <b>Mn</b> <small>54.94</small>	<small>26</small> <b>Fe</b> <small>55.85</small>	<small>27</small> <b>Co</b> <small>58.93</small>	<small>28</small> <b>Ni</b> <small>58.71</small>	<small>29</small> <b>Cu</b> <small>63.55</small>	<small>30</small> <b>Zn</b> <small>65.39</small>	<small>31</small> <b>Ga</b> <small>69.72</small>	<small>32</small> <b>Ge</b> <small>72.64</small>	<small>33</small> <b>As</b> <small>74.92</small>	<small>34</small> <b>Se</b> <small>78.96</small>	<small>35</small> <b>Br</b> <small>79.90</small>	<small>36</small> <b>Kr</b> <small>83.80</small>																															
<small>37</small> <b>Rb</b> <small>85.47</small>	<small>38</small> <b>Sr</b> <small>87.62</small>	<small>39</small> <b>Y</b> <small>88.91</small>	<small>40</small> <b>Zr</b> <small>91.22</small>	<small>41</small> <b>Nb</b> <small>92.91</small>	<small>42</small> <b>Mo</b> <small>95.94</small>	<small>43</small> <b>Tc</b> <small>98.00</small>	<small>44</small> <b>Ru</b> <small>101.07</small>	<small>45</small> <b>Rh</b> <small>101.07</small>	<small>46</small> <b>Pd</b> <small>106.36</small>	<small>47</small> <b>Ag</b> <small>107.87</small>	<small>48</small> <b>Cd</b> <small>112.41</small>	<small>49</small> <b>In</b> <small>114.82</small>	<small>50</small> <b>Sn</b> <small>118.71</small>	<small>51</small> <b>Sb</b> <small>121.76</small>	<small>52</small> <b>Te</b> <small>127.60</small>	<small>53</small> <b>I</b> <small>126.91</small>	<small>54</small> <b>Xe</b> <small>131.29</small>																															
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## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

## Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Section 2 nuclear Chemistry Study Guide Key

**Donna Batton**



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