

**Acids and Bases Worksheet**

NAME \_\_\_\_\_

1. Name the following compounds as acids and circle which are weak acids:



2. Write the formulas for the following bases:



3. Complete the following table:

	Acids	Bases
Definition		
Electrolyte in aqueous solution?		
pH		
Taste		
Litmus turns		
Which reacts with metals to form $\text{H}_2$ gas?		

**Conjugate acid-base pairs**

4. Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:



5. Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:



6. An acid will form the conjugate \_\_\_\_\_ and a base will form the conjugate \_\_\_\_\_

**Neutralization and Titration.**

7. Classify the following acids as monoprotic, diprotic or triprotic:

8. What is the concentration of an unknown  $\text{NaOH}$  solution if took an average of 15.4mL of 0.100N  $\text{HCl}$  solution to neutralize 10.0mL of the  $\text{NaOH}$  solution?9. From the given data, what is the molarity and the normality of the  $\text{HF}$  solution if the  $\text{KOH}$  solution was 0.50M?

	Acid, $\text{HF}$	Base, $\text{KOH}$
Initial Volume, mL	5.5	14.9
Final Volume, mL	25.9	24.4

**$\text{pH} = -\log[\text{H}_3\text{O}^+]$**  To find pH if given  $[\text{H}_3\text{O}^+]$ , enter the  $[\text{H}_3\text{O}^+]$ , press Log & then the +/- button  
To find  $[\text{H}_3\text{O}^+]$  if given pH, enter negative pH, then 2<sup>nd</sup> function Log

10. A solution has a  $[\text{H}_3\text{O}^+]$  of  $1 \times 10^{-7}$  M, what is its pH? \_\_\_\_ Is this an acid, base or is it neutral?

TURN THE PAPER OVER

11. What is the pH of a solution if the  $[\text{H}^+] = 7.2 \times 10^{-6}\text{M}$ ? \_\_\_\_ Is this an acid or base?

# Acids To Bases Answer Key

**Mary Millhollon, Richard H. Langley**



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