

### ACS Practice Test 1

#### Acids & Bases:

1. All are examples of Lewis acid-base reactions *except*
  - (A)  $\text{Cu}^{2+}(aq) + 4\text{NH}_3(aq) \rightleftharpoons [\text{Cu}(\text{NH}_3)_4]^{2+}(aq)$
  - (B)  $\text{HCl}(aq) + \text{NH}_3(aq) \rightarrow \text{NH}_4\text{Cl}(aq)$
  - (C)  $\text{H}^+(aq) + \text{OH}^-(aq) \rightleftharpoons \text{H}_2\text{O}(l)$
  - (D)  $2\text{Na}(s) + \text{Cl}_2(g) \rightarrow 2\text{NaCl}(s)$
2. According to the Lewis definition, an acid is a species
  - (A) having a hydrogen ion.
  - (B) donating a pair of electrons.
  - (C) accepting a pair of electrons.
  - (D) accepting a hydrogen ion.
3. All are potential Lewis bases *except*
  - (A)  $\text{NH}_3$
  - (B)  $\text{H}_2\text{O}$
  - (C)  $\text{CH}_4$
  - (D)  $\text{CN}^-$
4. Which hydrated metal ion is most acidic under conditions of equal molar concentration in water?
  - (A)  $\text{Al}^{3+}$
  - (B)  $\text{Ba}^{2+}$
  - (C)  $\text{K}^+$
  - (D)  $\text{Zn}^{2+}$
  - (E)  $\text{Ag}^+$
5. Which element in Group 5A forms the most basic oxide?
  - (A) N
  - (B) P
  - (C) As
  - (D) Sb
  - (E) Bi

#### Atomic Structure:

6. In what respect does an atom of magnesium differ from a magnesium ion ( $\text{Mg}^{2+}$ )?
  - (A) The ion has a more stable electronic arrangement than the atom.
  - (B) The positive charge on the nucleus of the ion is two units greater than the nuclear charge on the atom.
  - (C) The ion has two more protons than the atom.
  - (D) The ion will react more readily with free  $\text{Cl}_2$ .
  - (E) The ion has two more planetary electrons than the atom.
7. A certain atom has the symbol  $^{148}_{73}\text{Px}$ . What does this symbol tell about an atom of the element?
  - (A) It has 148 electrons.
  - (B) It has a valence of 4.
  - (C) It has 221 protons in the nucleus.
  - (D) It has 75 neutrons in the nucleus.
  - (E) Its atomic mass is 73 u.
8. The atomic mass of an element is 32.07 u and its atomic number is 16. The number of protons contained in the nucleus of an atom of this element is
  - (A) 16
  - (B) 31
  - (C) 32
  - (D) 48
9. The atomic mass of an element is 32.07 u and its atomic number is 16. The element forms a simple ion. The ion will most likely have a charge of
  - (A) 1-
  - (B) 2-
  - (C) 3+
  - (D) 1+

## Acs Practice Test 2 Answer Key

**Martha Humbard**

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**Field Manuals** United States. War Department,1948

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