

OTHS Pre-AP Chemistry Equations and Constants

The following symbols have the definitions specified unless otherwise noted.

g	= gram(s)
nm	= nanometer(s)
atm	= atmosphere(s)
L, mL	= liter(s), milliliter(s)

J, kJ	= joule(s), kilojoule(s)
mol	= mole(s)
mm Hg	= millimeters of mercury
Pa, kPa	= pascal(s), kilopascal(s)

ATOMS, ENERGY, & ELECTRONS

$$E_{\text{photon}} = h\nu$$

$$c = \lambda\nu$$

$$E_{\text{photon}} = \frac{hc}{\lambda}$$

$$\begin{aligned} E &= \text{energy} \\ \nu &= \text{frequency} \\ \lambda &= \text{wavelength} \end{aligned}$$

$$\begin{aligned} \text{Planck's constant, } h &= 6.626 \times 10^{-34} \text{ J} \cdot \text{s} \\ \text{Speed of light, } c &= 2.998 \times 10^8 \text{ m} \cdot \text{s}^{-1} \\ \text{Avogadro's number} &= 6.022 \times 10^{23} \text{ mol}^{-1} \end{aligned}$$

ACIDS and BASES

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

$$K_b = \frac{[OH^-][HB^+]}{[B]}$$

$$K_w = [H^+][OH^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$K_w = K_a \times K_b$$

$$pH = -\log[H^+], \quad pOH = -\log[OH^-]$$

$$[OH^-] = 10^{-pOH}, \quad [H^+] = 10^{-pH}$$

$$14 = pH + pOH$$

Equilibrium Constants

$$\begin{aligned} K_a &= \text{(weak acid)} \\ K_b &= \text{(weak base)} \\ K_{H_2O} &= \text{(water)} \end{aligned}$$

GASES, LIQUIDS, AND SOLUTIONS

$$PV = nRT$$

$$P_{\text{total}} = P_1 + P_2 + P_3 \dots$$

$$\frac{P_1 V_1}{T_1 n_1} = \frac{P_2 V_2}{T_2 n_2}$$

$$K = {}^\circ\text{C} + 273.15$$

$$n = \frac{m}{\text{molar mass}}$$

$$\text{Kinetic Energy} = \frac{1}{2}mv^2$$

$$\text{Molarity, } M = \frac{\text{moles solute}}{\text{liters solution}}$$

$$\text{Dilution formula, } M_1 V_1 = M_2 V_2$$

$$D = \frac{m}{V}$$

$$P = \text{pressure}$$

$$V = \text{volume}$$

$$T = \text{temperature}$$

$$n = \text{number of moles of gas}$$

$$\begin{aligned} R, \text{ Ideal Gas Constant} &= 0.08314 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \text{K}^{-1} \\ &= 0.08206 \text{ L} \cdot \text{atm} \cdot \text{mol}^{-1} \text{K}^{-1} \\ &= 62.36 \text{ L} \cdot \text{torr} \cdot \text{mol}^{-1} \text{K}^{-1} \\ &= 8.314 \text{ J} \cdot \text{mol}^{-1} \text{K}^{-1} \end{aligned}$$

$$K = \text{Kelvin temperature}$$

$${}^\circ\text{C} = \text{Celsius temperature}$$

$$1 \text{ atm} = 760.0 \text{ mm Hg} = 760.0 \text{ torr} = 101.3 \text{ kPa}$$

$$STP = 273.15 \text{ K and } 1.000 \text{ atm}$$

$$\text{Volume of an Ideal Gas at STP} = 22.40 \text{ L} \cdot \text{mol}^{-1}$$

$$v = \text{velocity}$$

$$D = \text{density}, m = \text{mass}$$

$$1 \text{ cm}^3 = 1 \text{ mL}$$

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